

Introduction 11

An Electron's Address

Name: _____
 Date: _____
 Class: _____

Introduction: Energy Levels and Sublevels

The professor in his laboratory would have prepared the following labeled energy levels. The electron would either occur just as energy in these sublevels.

Remember that labeled energy levels just like sublevels labeled levels occur. Each sublevel occurs once. Draw the following table:

Table 1

Energy Level	Sublevels in energy level and their corresponding orbitals
1	s
2	s, p
3	s, p, d
4	s, p, d, f

Now that there is a table draw as a structural model of the 4th energy level because there are 4 levels of sublevels in the 4th energy level.

Critical Thinking Questions

- Which energy sublevel occurs in the 1st energy level?
 The sublevel that occurs in the 1st energy level is s and sublevel is the only one sublevel.
- Which energy sublevel occurs in the 2nd energy level?
 The sublevels that occur in the 2nd energy level is s, p, and sublevel.
- Which energy sublevel occurs in the 3rd energy level?
 The sublevels that occur in the 3rd energy level are s, p, d, and sublevel.
- Which energy sublevel would you expect to find in the 4th energy level?
 The number of sublevels I would expect to find in the 4th energy level would be 4 sublevels. There would be 4 sublevels and the number of sublevels equals the energy level number.
- Which is "a sublevel?" (There are 4th levels for the 4th energy level.)
 The 4th energy level would have 4 sublevels (s, p, d, and f). The 4th energy level would be the 4th energy level (s, p, d, and f).

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