

Part 1: Multiple Choice

1. A solution of 0.10 M NaCl and 0.10 M Na₂CO₃ is titrated with 0.10 M HCl. The pH of the solution is 10.00 at the start of the titration. The pH of the solution is 7.00 at the equivalence point. The pH of the solution is 4.00 at the end of the titration. The pH of the solution is 1.00 at the end of the titration.

2. A solution of 0.10 M NaCl and 0.10 M Na₂CO₃ is titrated with 0.10 M HCl. The pH of the solution is 10.00 at the start of the titration. The pH of the solution is 7.00 at the equivalence point. The pH of the solution is 4.00 at the end of the titration. The pH of the solution is 1.00 at the end of the titration.

3. A solution of 0.10 M NaCl and 0.10 M Na₂CO₃ is titrated with 0.10 M HCl. The pH of the solution is 10.00 at the start of the titration. The pH of the solution is 7.00 at the equivalence point. The pH of the solution is 4.00 at the end of the titration. The pH of the solution is 1.00 at the end of the titration.

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5. A solution of 0.10 M NaCl and 0.10 M Na₂CO₃ is titrated with 0.10 M HCl. The pH of the solution is 10.00 at the start of the titration. The pH of the solution is 7.00 at the equivalence point. The pH of the solution is 4.00 at the end of the titration. The pH of the solution is 1.00 at the end of the titration.

6. A solution of 0.10 M NaCl and 0.10 M Na₂CO₃ is titrated with 0.10 M HCl. The pH of the solution is 10.00 at the start of the titration. The pH of the solution is 7.00 at the equivalence point. The pH of the solution is 4.00 at the end of the titration. The pH of the solution is 1.00 at the end of the titration.

- a. _____
- b. _____
- c. _____
- d. _____
- e. _____

7. A solution of 0.10 M NaCl and 0.10 M Na₂CO₃ is titrated with 0.10 M HCl. The pH of the solution is 10.00 at the start of the titration. The pH of the solution is 7.00 at the equivalence point. The pH of the solution is 4.00 at the end of the titration. The pH of the solution is 1.00 at the end of the titration.

8. A solution of 0.10 M NaCl and 0.10 M Na₂CO₃ is titrated with 0.10 M HCl. The pH of the solution is 10.00 at the start of the titration. The pH of the solution is 7.00 at the equivalence point. The pH of the solution is 4.00 at the end of the titration. The pH of the solution is 1.00 at the end of the titration.

9. A solution of 0.10 M NaCl and 0.10 M Na₂CO₃ is titrated with 0.10 M HCl. The pH of the solution is 10.00 at the start of the titration. The pH of the solution is 7.00 at the equivalence point. The pH of the solution is 4.00 at the end of the titration. The pH of the solution is 1.00 at the end of the titration.

10. A solution of 0.10 M NaCl and 0.10 M Na₂CO₃ is titrated with 0.10 M HCl. The pH of the solution is 10.00 at the start of the titration. The pH of the solution is 7.00 at the equivalence point. The pH of the solution is 4.00 at the end of the titration. The pH of the solution is 1.00 at the end of the titration.



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