

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
 potassium-40 undergoes beta decay (electron)  
 uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.  
 Write the electron configuration for Cr.  
 Write the shorthand notation for Bi.



### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- |                             |   |
|-----------------------------|---|
| a. Sulfur hexafluoride      | $\text{SF}_6$   |
| b. Lithium nitride          | $\text{Li}_3\text{N}$   |
| c. Chromium (III) carbonate | $\text{Cr}_2(\text{CO}_3)_3$  |
| d. Tin (II) chloride        | $\text{SnCl}_2$   |
| e. Ammonium acetate         | $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$ ) |
| f. Mercury (I) chloride     | $\text{Hg}_2\text{Cl}_2$  |
| g. Potassium bromate        | $\text{KBrO}_3$   |
| h. Hydrosulfuric acid       | $\text{H}_2\text{S(aq)}$  |
| i. Chloric acid             | $\text{HClO}_3$   |
| j. Sulfurous acid           | $\text{H}_2\text{SO}_3$   |

Name each of the following compounds:

- |                              |                         |
|------------------------------|-------------------------|
| a. $\text{CuSO}_4$           | Copper (II) sulfate     |
| b. $\text{AlF}_3$            | aluminum fluoride       |
| c. HI                        | hydrogen iodide         |
| d. NO                        | nitrogen monoxide       |
| e. $\text{H}_2\text{Se(aq)}$ | hydroselenic acid       |
| f. $\text{HNO}_3$            | nitric acid             |
| g. $\text{HNO}_2$            | nitrous acid            |
| h. $\text{NaHSO}_4$          | sodium hydrogen sulfate |

### Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



# Modern Chemistry Chapter 5 Mixed Review Answers

**SA Adler**



## Modern Chemistry Chapter 5 Mixed Review Answers:

**Modern Chemistry** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001     *Modern Machinery*, 1907     I & EC, 1925     *Industrial & Engineering Chemistry*, 1925     **Saturday Review**, 1876     **Modern Electrochemistry 2B**  
John O'M. Bockris, Amulya K.N. Reddy, 2007-05-08 This book had its nucleus in some lectures given by one of us J O M B in a course on electrochemistry to students of energy conversion at the University of Pennsylvania It was there that he met a number of people trained in chemistry physics biology metallurgy and materials science all of whom wanted to know something about electrochemistry The concept of writing a book about electrochemistry which could be understood by people with very varied backgrounds was thereby engendered The lectures were recorded and written up by Dr Klaus Muller as a 293 page manuscript At a later stage A K N R joined the effort it was decided to make a fresh start and to write a much more comprehensive text Of methods for direct energy conversion the electrochemical one is the most advanced and seems the most likely to become of considerable practical importance Thus conversion to electrochemically powered transportation systems appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing concentration in the atmosphere of carbon dioxide may be met Cor sion is recognized as having an electrochemical basis The synthesis of nylon now contains an important electrochemical stage Some central biological mechanisms have been shown to take place by means of electrochemical reactions A number of American organizations have recently recommended greatly increased activity in training and research in electrochemistry at universities in the United States     Practical Chemistry  
Newton Henry Black, James Bryant Conant, 1920     **The Saturday Review of Politics, Literature, Science and Art**, 1876  
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**Scientific American**, 1898     Journal of Gas Lighting, 1917     **Book Review Digest**, 1996-08     New Scientist, 1982  
*The Athenaeum*, 1876     *Choice*, 2001     Journal of Education and School World, 1889

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In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Modern Chemistry Chapter 5 Mixed Review Answers**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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