

Half Modern Chemistry Review

CHAPTER 9: STOICHIOMETRY

The following pages contain the best (but not all) of the information for the chapter 9 test.

Focus on this content, but make sure to review class notes, activities, worksheets, questions, etc.

If you study this document and NOTHING else, you should at least be able to PASS the test.

**** Test items will be recall, recognition, and/or application of this content. ****

OBJECTIVES

- Calculate molar masses for molecular-formula compounds. (C & C)
- Calculate molar masses for ionic compounds and gases. (B & C)
- 9.1. Determine masses of reactants from balanced chemical equations. (T & P)
- 9.2. Perform stoichiometry calculations with: mol to mol, mol to gram, gram to mol, and gram to gram. (P & P)
- 9.3. Calculate stoichiometry problems and determine the limiting reagent, excess reactant, and amount of product produced. (P & P)

9.1: INTRODUCTION TO STOICHIOMETRY

a. Introductory

- **stoichiometry** = stoichiometry = calculations involving the mass relationships of elements in compounds.
- **molar mass** = molar mass = calculations involving the mass relationships between reactants and products in a chemical reaction.
- **molar ratio** = a conversion factor that relates the amount, in moles, of any two substances involved in a chemical reaction.

b. Chapter Highlights

- Relates stoichiometry to the mass relationships between reactants and products in a chemical reaction.
- Stoichiometry is based on the law of conservation of mass.
- Relating one substance to another is called a **stoichiometric conversion** of mass.
- A **molar ratio** is the conversion factor that relates the amount, in moles, of any two substances in a chemical reaction. This molar ratio is derived from the balanced equation.
 - Example: $2\text{H}_2\text{O} + \text{BCl}_3 \rightarrow \text{B}_2\text{O}_3 + 2\text{H}_2\text{O}$ (What is the ratio of mass of oxygen to water? 4:16 or 1:4)
- Amount of a substance is measured in moles, and mass of a substance is represented using mass units such as grams, kilograms, or milligrams.
- Mass and amount of substances are quantities, whereas moles and grams are units.
- A balanced chemical equation is necessary to solve any stoichiometric problem.

9.2: MASS STOICHIOMETRIC CALCULATIONS

a. Chapter Highlights

- We use molar stoichiometric calculations, the mass, or the amount of any reactant or product over the calculation of the balanced chemical equation, and the mass or amount of any other reactant or product is known.

9.3: LIMITED REACTANTS AND PERCENTAGE YIELD

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Richard Bailey



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