

Name : \_\_\_\_\_

Date : \_\_\_\_\_

## MOLARITY AND DILUTION WORKSHEET

1. 40.0 mL of 0.400 M NaOH solution is diluted to a final volume of 200.0 mL. Calculate the new concentration.
2. A solution of 85.0 mL of NaOH is diluted to a final volume of 290.0 mL, and the new molarity is 0.0500 M. Calculate the original molarity of the base.
3. 150.0 mL of 0.025 M NaOH solution is added to 150.0 mL of water. Calculate the new molarity.
4. 220.0 mL of a solution of NaOH is diluted by adding 250.0 mL of water to produce a new molarity of 0.0500 M. Calculate the molarity of the base.
5. A 0.350 M solution is concentrated by evaporation to a reduced final volume of 100.0 mL and molarity of 0.825 M. Calculate the original volume.
6. 850.0 mL of 0.280 M KOH solution is diluted to a final volume of 1000.0 mL. Calculate the new concentration.

# Molarity By Dilution 69 Answers

**JG Myers**



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