



### 3: Chemical Kinetics

Standard: 12th Science

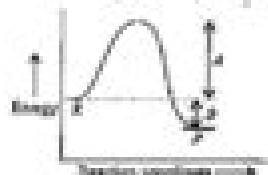
Subject: Chemistry

Student Name: \_\_\_\_\_

**Q1.** Consider the graph given in Fig. Which of the following options does not show instantaneous rate of reaction at 40<sup>th</sup> second?

1.  $\frac{V_2 - V_1}{50 - 30}$
2.  $\frac{V_1 - V_2}{50 - 30}$
3.  $\frac{V_2 - V_1}{40 - 30}$
4.  $\frac{V_1 - V_2}{40 - 30}$

**Q2.** The energy diagram of a reaction  $P + Q \rightarrow R + S$  is given. What are A and B in the graph?



1. A → activation energy, B → heat of reaction
  2. A → threshold energy, B → heat of reaction
  3. A → heat of energy, B → activation reaction
  4. A → potential energy, B → energy of reaction
- Q3.** Decomposition of nitrogen pentoxide is known to be a first order reaction 75 percent of the oxide had decomposed in the first 24 minutes. At the end of an hour, after the start of the reaction, the amount of oxide left will be
1. Nil
  2. About 1%
  3. About 2%
  4. About 3%
- Q4.** Higher order (>3) reactions are rare due to:
1. Low probability of simultaneous collision of all the reacting species.
  2. Increase in entropy and activation energy as more molecules are involved.



# Practice Chemical Kinetics Questions Answer

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