

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
 potassium-40 undergoes beta decay (electron)  
 uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.  
 Write the electron configuration for Cr.  
 Write the shorthand notation for Bi.



### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds:

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



# Chapter 7 Modern Chemistry Answers

**RJ Shavelson**



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The Chautauquan ,1889 TEXT BOOK OF MODERN PHARMACEUTICAL ANALYTICAL TECHNIQUES Valapa Anusha, Dr. Deepak Shrivastava, Ms. S. Thasneem, Dr. Pushpendar Bardiya, Dr. Khushboo Mishra,2025-09-20 The Textbook of Modern Pharmaceutical Analytical Techniques is a comprehensive guide that explores a wide range of analytical tools essential for pharmaceutical sciences It begins with UV Visible spectroscopy covering its introduction theoretical principles governing laws instrumentation solvent effects and diverse applications in drug analysis The book then moves into Infrared IR spectroscopy explaining molecular vibrations sample handling dispersive and Fourier Transform IR spectrometers factors influencing vibrational frequencies and its significance in pharmaceutical applications A detailed chapter on Spectrofluorimetry highlights the theory of fluorescence influencing factors quenchers instrumentation and its vital role in qualitative and quantitative analysis Further Flame Emission Spectroscopy FES and Atomic Absorption Spectroscopy AAS are thoroughly explained focusing on principles instrumentation interferences and pharmaceutical applications especially in trace metal analysis The text also covers Nuclear Magnetic Resonance NMR spectroscopy providing insights into quantum numbers basic principles instrumentation solvent requirements relaxation processes signal interpretation chemical shifts spin spin coupling coupling constants and advanced techniques like FT NMR and <sup>13</sup>C NMR The applications of NMR in structural elucidation of drugs are given special emphasis Following this Mass Spectroscopy is presented with clarity elaborating its principle instrumentation ionization techniques EI CI FAB MALDI ESI APCI APPI types of analyzers fragmentation rules metastable ions isotopic peaks and wide ranging pharmaceutical applications A large portion of the book is devoted to Chromatography offering a complete discussion on principles apparatus instrumentation chromatographic parameters and factors affecting resolution across various techniques These include paper chromatography thin layer chromatography TLC ion exchange chromatography column chromatography gas chromatography GC high performance liquid chromatography HPLC and affinity chromatography Each method is explained with its specific advantages and pharmaceutical uses The section on Electrophoresis elaborates on different types such as paper gel capillary zone moving boundary and isoelectric focusing describing their principles instrumentation working conditions influencing factors and applications in protein and drug separation The book also introduces X ray Crystallography explaining X ray production diffraction methods Bragg s law rotating crystal technique X ray powder diffraction crystal types and applications in

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