

# 8

## Acids, bases and salts

### Key objectives

By the end of this section, you should be able to:

- describe the characteristic reactions of acids
  - state that bases are oxides or hydroxides of metals
  - state that alkalis are soluble bases
  - describe the characteristic reactions of bases
  - state that aqueous solutions of acids contain  $\text{H}^+$  ions and aqueous solutions of alkalis contain  $\text{OH}^-$  ions
  - describe how to use pH, as measured with universal indicator paper, to compare hydrogen ion concentration, neutrality, relative acidity and alkalinity
  - describe the neutralisation reaction
- define acids and bases in terms of proton transfer
  - explain the difference between weak and strong acids in terms of dissociation
- classify oxides as acidic, basic or amphoteric, with examples
  - describe the general solubility rules for salts
  - describe the preparation of soluble salts by the reaction of an acid with:
    - an alkali
    - excess metal, insoluble base or carbonate
- describe the preparation of insoluble salts by precipitation
- define a hydrated substance and an anhydrous substance
- define the term water of crystallisation in crystals, including  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  and  $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$

### Key terms

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Term	Definition
<b>Acid</b>	A substance which dissolves in water, producing $\text{H}^+(\text{aq})$ ions as the only positive ion. <b>A proton (<math>\text{H}^+</math>) donor.</b>
<b>Alkali</b>	A soluble base which produces $\text{OH}^-(\text{aq})$ ions in water.
<b>Anhydrous salt</b>	A salt which has had its water of crystallisation removed.
<b>Base</b>	A substance which neutralises an acid, producing a salt and water as the only products. Bases are oxides or hydroxides of metals. (Ammonia is also a base.) <b>A proton (<math>\text{H}^+</math>) acceptor.</b>
<b>Indicator</b>	A substance that shows whether a substance is acidic or alkaline by changing colour.
<b>Neutralisation</b>	The process in which an acid reacts with a base to form water.
<b>pH scale</b>	A scale running from 0 to 14 used to express the acidity or alkalinity of a substance.
<b>Saturated solution</b>	A solution containing the maximum amount of dissolved solute in the solvent at a given temperature.
<b>Water of crystallisation</b>	Water incorporated into the structure of a substance as it crystallises, for example in copper(II) sulfate pentahydrate ( $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ ).

### 8.1 Acids and alkalis

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#### Acids

**Acids** are substances that produce  $\text{H}^+$  ions when they are dissolved in water.

Acids are defined as proton ( $\text{H}^+$ ) donors.

# Chapter 8 Acids Bases And Ph Quia

**Terry C. Jones**



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