

Chemical Reaction Engineering

The rate of a reaction r_A

Since chemical reaction engineering concerns itself with the kinetics of the reaction, we must first define what is meant by the "rate of a reaction".

We define the rate of a reaction for component A as follows: "the amount of moles of A disappearing or being generated per unit time per unit volume is the rate of the reaction, denoted as r_A ".

Equation for Rate

The rate of a reaction is determined by the rate law. For example, for the reaction,



The rate law could look like,

$$-r_A = k[A]$$

Here, $[A]$ means the concentration of A in mol/L. We will use C_A to denote the concentration from now on. So,

$$-r_A = kC_A$$

In this case we say that the reaction is **first order in A**, as the exponent of C_A is 1. The rate law could also look like,

$$-r_A = kC_A^2$$

In this case, the reaction would be second order in A.

The rate law could also look something like this,

$$-r_A = \frac{k_1 C_A}{1 + k_2 C_A}$$

In which case the order would be hard to determine.

Rate for components of a reaction

Given the reaction,



The rate laws for each component are related as,

$$\frac{1}{a} r_A = \frac{1}{b} r_B = \frac{1}{c} r_C = \frac{1}{d} r_D$$

General Mole Balance for a System / Reactor

We now consider a system volume with a species j as shown below:

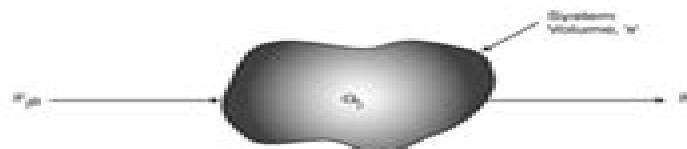


Figure 1-5 : Mole balance on species j in a system volume, V

The mole balance on this system is,

$$\begin{aligned} & \text{[moles of j in per unit time]} - \text{[moles of j out per unit time]} + \text{[moles of j generated per unit time]} \\ & = \text{[moles of j accumulated per unit time]} \end{aligned}$$

$$F_{j0} - F_j + G_j = \frac{dN_j}{dt}$$

If all the properties of a system, such as catalysis concentration, temperature, concentration of species are uniform throughout the volume, then we can write the G_j terms simply as the product of r_j and the system volume V.

$$F_{j0} - F_j + r_j V = \frac{dN_j}{dt}$$

Chemical Reaction Engineering Notes

Peipei Pang



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Table of Contents Chemical Reaction Engineering Notes

1. Understanding the eBook Chemical Reaction Engineering Notes
 - The Rise of Digital Reading Chemical Reaction Engineering Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Reaction Engineering Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Reaction Engineering Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Reaction Engineering Notes
 - Personalized Recommendations
 - Chemical Reaction Engineering Notes User Reviews and Ratings
 - Chemical Reaction Engineering Notes and Bestseller Lists
5. Accessing Chemical Reaction Engineering Notes Free and Paid eBooks
 - Chemical Reaction Engineering Notes Public Domain eBooks
 - Chemical Reaction Engineering Notes eBook Subscription Services

- Chemical Reaction Engineering Notes Budget-Friendly Options
- 6. Navigating Chemical Reaction Engineering Notes eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Reaction Engineering Notes Compatibility with Devices
 - Chemical Reaction Engineering Notes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Reaction Engineering Notes
 - Highlighting and Note-Taking Chemical Reaction Engineering Notes
 - Interactive Elements Chemical Reaction Engineering Notes
- 8. Staying Engaged with Chemical Reaction Engineering Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Reaction Engineering Notes
- 9. Balancing eBooks and Physical Books Chemical Reaction Engineering Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Reaction Engineering Notes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Reaction Engineering Notes
 - Setting Reading Goals Chemical Reaction Engineering Notes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Reaction Engineering Notes
 - Fact-Checking eBook Content of Chemical Reaction Engineering Notes
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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