

# KINETICS Practice Problems and Solutions

Name: \_\_\_\_\_  
 Period: \_\_\_\_\_ Date: \_\_\_\_\_

AP Chemistry  
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
- This increases the  $[H_2]$ , which will increase the rate, but has no effect on  $k$
  - Due to Arrhenius equation, changing temperature changes the value of  $k$ .
  - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of  $k$ .

6. Consider the reaction:  $P_4 + 6 H_2 \rightarrow 4 PH_3$ . A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$ , mol/L	$[H_2]$ , mol/L	Initial Rate, mol/L · s
0.0110	0.0075	$3.20 \times 10^{-4}$
0.0110	0.0150	$6.40 \times 10^{-4}$
0.0220	0.0150	$6.39 \times 10^{-4}$

- What is the order with respect to:  $P_4$  0  
 $H_2$  1
- Write the rate law for this reaction.  $rate = k[H_2]$
- Determine the value and units of the rate constant,  $k$ . **plug and chug using the rate law & data from exp't 1 and solving for  $k$ , we get  $k = 0.0427 s^{-1}$**

7. Consider the reaction:  $SO_2 + O_3 \rightarrow SO_3 + O_2$ . A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$ , mol/L	$[O_3]$ , mol/L	Initial Rate, mol/L · s
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- What is the order with respect to:  $SO_2$  2  
 $O_3$  0
- Write the rate law for this reaction.  $rate = k[SO_2]^2[O_3]^0$
- Determine the value and units of the rate constant,  $k$ . **plug and chug using the rate law & data from exp't 1 and solving for  $k$ , we get  $k = 2.36 molL^{-2} s^{-1}$**

8. Consider the following mechanism.
- $$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$
- $$A_2 + R \rightarrow C \quad (\text{fast})$$

- Write the overall balanced chemical equation.  $2 A_2 + B_2 \rightarrow 2 C$
- Identify any intermediates within the mechanism. **R**
- What is the order with respect to each reactant?  **$A_2$  1<sup>st</sup>;  $B_2$  1<sup>st</sup>**

# Chemical Kinetics Practice Problems And Solutions

**Oswaal Editorial Board**



## **Chemical Kinetics Practice Problems And Solutions :**

The book delves into Chemical Kinetics Practice Problems And Solutions . Chemical Kinetics Practice Problems And Solutions is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chemical Kinetics Practice Problems And Solutions , encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to Chemical Kinetics Practice Problems And Solutions
- Chapter 2: Essential Elements of Chemical Kinetics Practice Problems And Solutions
- Chapter 3: Chemical Kinetics Practice Problems And Solutions in Everyday Life
- Chapter 4: Chemical Kinetics Practice Problems And Solutions in Specific Contexts
- Chapter 5: Conclusion

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