

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

## States of Matter

## Section 13.1 Gases

In your textbook, read about the kinetic molecular theory.

Complete each statement.

- The kinetic molecular theory describes the behavior of gases in terms of particles in motion.
- The kinetic molecular theory makes the following assumptions:
  - In a sample of a gas, the volume of the gas particles themselves is very small compared to the volume of the sample.
  - Between gas particles and the walls, there are no significant attractive or repulsive forces between gas particles.
  - Gas particles are in constant and random motion.
  - The collisions between gas particles are elastic, that is, no kinetic energy is lost.
- The kinetic energy of a particle is represented by the equation  $KE = \frac{1}{2}mv^2$ .
- Temperature is a measure of the average kinetic energy of the particles in a sample of matter.

In your textbook, read about explaining the behavior of gases.

For each statement below, write true or false.

- |              |  |
|--------------|--|
| <u>True</u>  | 1. Gases are less dense than solids because there is a lot of space between the particles of a gas.                              |
| <u>True</u>  | 2. The random motion of gas particles causes a gas to expand until it fills its container.                                       |
| <u>False</u> | 3. The density of a gas decreases as it is compressed.   |
| <u>True</u>  | 4. A gas can flow into a space occupied by another gas.  |
| <u>True</u>  | 5. The diffusion of a gas is caused by the random motion of the particles of the gas.  |
| <u>False</u> | 6. Lighter gas particles diffuse less rapidly than do heavier gas particles.   |
| <u>True</u>  | 7. During effusion, a gas escapes through a tiny opening into a vacuum.  |
| <u>False</u> | 8. Graham's law of effusion states that the rate of effusion for a gas is directly related to the square root of its molar mass. |

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## CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

## Section 13.1 continued

In your textbook, read about gas pressure.

Circle the letter of the choice that best completes the statement or answers the question.

- Pressure is defined as force per unit
  - area
  - time
  - mass
  - volume
- What is an instrument designed to measure atmospheric pressure?
  - barometer
  - manometer
  - physoconimeter
  - densimeter
- The height of the liquid in a barometer is affected by all of the following EXCEPT the
  - altitude
  - density of the liquid in the column
  - atmospheric pressure
  - diameter of the column tube.
- The pressure of the gas in a manometer is directly related to which of the following
  - height of the mercury column in the closed-end arm
  - height of the mercury column in the open-end arm
  - $h + 76$
  - $h - 76$
- One atmosphere is equal to a pressure of
  - 76 mm Hg
  - 101.3 kPa
  - 147 psi
  - 760 mm
- The partial pressure of a gas depends on all of the following EXCEPT the
  - concentration of the gas
  - size of the container
  - density of the gas
  - temperature of the gas
- The pressure of a sample of air in a manometer is 101.3 kPa. What is the partial pressure of nitrogen ( $N_2$ ) in the sample if the combined partial pressures of the other gases is 22.4 kPa?
  - 42.4 kPa
  - 78.9 kPa
  - 101.3 kPa
  - 123.7 kPa

Use the figure to answer the following questions.

- What instrument is illustrated in the figure? barometer
- Who invented this instrument? Evangelista Torricelli
- What are the two opposing forces that control the height of the mercury in the column? gravity and atmospheric pressure
- What does it mean when the level of mercury drops in the column? The atmospheric pressure has increased.



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# Chemistry Matter And Change Chapter 11 Study Guide

**Sebastian Brünink**



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exam fee 51 15 other information you can take this test on computer kindly note that this is not an open book test reschedule 2 days in advance by 5 00 pm pacific standard time cancel 2 days in advance by 5 00 pm pacific standard time schedule 1 day in advance by 4 00 pm pacific standard time

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the california bar has three primary tasks the last task involves the oversight of 6 397 smog check stations in california as well as 1 577 brake and lamp check stations the stations must be licensed through the bar to provide smog checks if a station violates the laws then the bar can step in and investigate to possibly revoke the station

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this free online quiz will test your ability of understanding what to do when conducting a smog inspection here are some knowledge base questions when the sample hose doesn't reach the vehicle what is the procedure gas calibration when is the calibration performed what is hc hang up what can cause this and why

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class of license select the license s for which you are applying smog check inspector eo may inspect and certify the emissions control systems on vehicles subject to the smog check program at a licensed smog check test only or test and repair station

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oct 16 2023 california smog automotive institute 3001 chicago ave riverside ca 92507 united states google map phone 9517810412 view venue website bar update repair technicians smog inspector level 2 class schedule october 16 2023 to december 13 2023 6 00 pm to 10 00 pm monday tuesday wednesday thursday this

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tsi before starting the asm 50 15 test on a automatic transmission vehicle what is the correct gear that the inspector should

select default drive during a smog inspection you notice that the coolant radiator has a hole on it how should the inspector proceed with the smog test do not continue with the test