

Chemistry - Chapter 7 Study Guide

Section 7.1:

- *Valence electrons* – electrons in the outer energy level
 - responsible for bonding with other atoms
- How do you know how many valence electrons an element has?
 - Look at the group number
 - Group 1A = 1 valence electron
 - Group 7A = 7 valence electrons
 - Etc...
- *Electron Dot Diagrams* – a way of showing and keeping track of valence electrons
 - Written using the following steps:
 1. Write the element symbol – representing the nucleus and inner electrons
 2. Put one dot for each valence electrons,
 3. Don't forget to add them using Hund's rule – they don't pair up unless they have to.
- *The Octet Rule* – atoms tend to want to achieve a noble gas configuration because it is the most stable. Octet = eight electrons
- *Cations* – a positive ion (usually metals)
 - Lose electrons to attain a noble gas configuration
 - Written like Na^{1+}
 - No dots are shown (on an electron dot diagram) because they lose the valence electrons they had
- *Anions* – a negative ion (usually non-metals)
 - Gain electrons to attain a noble gas configuration
 - Written like P^{3-}
 - Dots are shown because valence electrons are gained
- 8 valence electrons = stable atom (noble gases.)

Section 7.2:

- Anions (-) and cations (+) are held together by opposite charges; a bond is formed through the transfer of electrons to achieve a noble gas configuration (octet.)
- *Ionic Compounds:*
 - Also called salts
 - Made from a cation and an anion
 - Which means they're made of a metal and a non-metal
- ***if you don't understand how to do ionic bonding, check the book or ask somebody to help... it's hard to explain***
- The result of ionic bonding is an ionic compound, shown in a *formula unit*
 - Written like Ca_3P_2

Chemistry Matter And Change Chapter 7 Study Guide Answer Key

Padhraic Smyth



Chemistry Matter And Change Chapter 7 Study Guide Answer Key:

The book delves into Chemistry Matter And Change Chapter 7 Study Guide Answer Key. Chemistry Matter And Change Chapter 7 Study Guide Answer Key is a crucial topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Chemistry Matter And Change Chapter 7 Study Guide Answer Key, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Chapter 2: Essential Elements of Chemistry Matter And Change Chapter 7 Study Guide Answer Key
 - Chapter 3: Chemistry Matter And Change Chapter 7 Study Guide Answer Key in Everyday Life
 - Chapter 4: Chemistry Matter And Change Chapter 7 Study Guide Answer Key in Specific Contexts
 - Chapter 5: Conclusion
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 3. In chapter 2, this book will delve into the foundational concepts of Chemistry Matter And Change Chapter 7 Study Guide Answer Key. The second chapter will elucidate the essential principles that must be understood to grasp Chemistry Matter And Change Chapter 7 Study Guide Answer Key in its entirety.
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 5. In chapter 4, the author will scrutinize the relevance of Chemistry Matter And Change Chapter 7 Study Guide Answer Key in specific contexts. This chapter will explore how Chemistry Matter And Change Chapter 7 Study Guide Answer Key is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, the author will draw a conclusion about Chemistry Matter And Change Chapter 7 Study Guide Answer Key. This chapter will summarize the key points that have been discussed throughout the book.
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