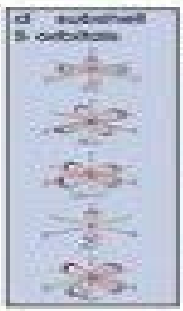


# CHEMISTRY

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## PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box indicating its position: **1.008** - atomic weight, **H** - atomic symbol, **1** - atomic number.



### ATOMIC STRUCTURE

Atomic Number,  $Z$ , is of protons in the nucleus. For neutral atom,  $Z = +$  of electrons. For ions of charge,  $+Q$ ,  $Z = +Q$  (cations) or  $Z = -Q$  (anions) of electrons in the nucleus. Atomic weight,  $A$ , is  $Z + N$ , the number of nucleons in the nucleus. Isotopes of atoms with the same  $Z$ , different  $N$ , are denoted by their atomic weight. For example,  $^{12}\text{C}$  and  $^{13}\text{C}$  are isotopes of carbon. The atomic weight is a weighted average of isotopes for a given element.

### ATOMIC QUANTUM NUMBERS AND ORBITALS

**Principal quantum number,  $n$** :  $n = 1, 2, 3, \dots$   
**Angular momentum quantum number,  $l$** :  $l = 0, 1, 2, \dots, n-1$   
**Magnetic quantum number,  $m_l$** :  $m_l = -l, -(l-1), \dots, 0, \dots, (l-1), l$   
**Spin quantum number,  $m_s$** :  $m_s = +\frac{1}{2}, -\frac{1}{2}$

### NUCLEAR CHEMISTRY

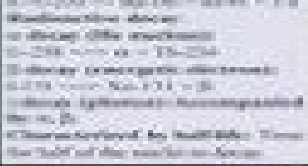
Nuclear reactions alter the number, particles, mass, and energy of atoms. Nuclear reactions are not affected by chemical reactions. Energy is released in nuclear reactions. Nuclear reactions are denoted by  ${}^A_Z\text{X}$ , where  $A$  is mass number,  $Z$  is atomic number, and  $\text{X}$  is the element symbol. Nuclear reactions are denoted by  ${}^A_Z\text{X} \rightarrow {}^A_Z\text{Y} + {}^A_Z\text{Z}$ .

### MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the behavior of electrons in atoms. The Pauli Exclusion Principle states that no two electrons in an atom can have the same set of quantum numbers. The Aufbau Principle states that electrons fill orbitals in order of increasing energy.

$n$	$l$	$m_l$	$m_s$	Orbitals
1	0	0	$+\frac{1}{2}, -\frac{1}{2}$	1s
2	0	0	$+\frac{1}{2}, -\frac{1}{2}$	2s
2	1	0	$+\frac{1}{2}, -\frac{1}{2}$	2p <sub>z</sub>
2	1	$\pm 1$	$+\frac{1}{2}, -\frac{1}{2}$	2p <sub>x</sub> , 2p <sub>y</sub>
3	0	0	$+\frac{1}{2}, -\frac{1}{2}$	3s
3	1	0	$+\frac{1}{2}, -\frac{1}{2}$	3p <sub>z</sub>
3	1	$\pm 1$	$+\frac{1}{2}, -\frac{1}{2}$	3p <sub>x</sub> , 3p <sub>y</sub>
3	2	0	$+\frac{1}{2}, -\frac{1}{2}$	3d <sub>z<sup>2</sup></sub>
3	2	$\pm 1$	$+\frac{1}{2}, -\frac{1}{2}$	3d <sub>xy</sub> , 3d <sub>xz</sub> , 3d <sub>yz</sub>
3	2	$\pm 2$	$+\frac{1}{2}, -\frac{1}{2}$	3d <sub>x<sup>2</sup>-y<sup>2</sup></sub>

**Atomic configuration:** Shows the arrangement of electrons in an atom. The Pauli Exclusion Principle states that no two electrons in an atom can have the same set of quantum numbers. The Aufbau Principle states that electrons fill orbitals in order of increasing energy. The order of filling is given by the Aufbau Principle.



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**Clemens Wendtner**



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