

**CHAPTER 4 PRACTICE TEST: SOLUTION STOICHIOMETRY**

1) Which of the following are weak electrolytes? HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$ , KCl

A) HCl, KCl

B) HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$ , KCl

C)  $\text{HC}_2\text{H}_3\text{O}_2$ , KCl

D)  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$

E) HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ , KCl

Answer: D Diff: 2

2) What are the spectator ions in the reaction between KOH (aq) and  $\text{HNO}_3$ (aq)?

A)  $\text{K}^+$  and  $\text{H}^+$

C)  $\text{K}^+$  and  $\text{NO}_3^-$

E)  $\text{OH}^-$  only

B)  $\text{H}^+$  and  $\text{OH}^-$

D)  $\text{H}^+$  and  $\text{NO}_3^-$

Answer: C Diff: 2

3) The balanced net ionic equation for precipitation of  $\text{CaCO}_3$  when aqueous solutions of  $\text{Na}_2\text{CO}_3$  and  $\text{CaCl}_2$  are mixed is \_\_\_\_\_.

A)  $2\text{Na}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Na}_2\text{CO}_3(\text{aq})$

E)

B)  $2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq})$

$\text{Na}_2\text{CO}_3(\text{aq}) + \text{CaCl}_2(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq}) + \text{CaCO}_3(\text{s})$

C)  $\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{NaCl}(\text{aq})$

D)  $\text{Ca}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{CaCO}_3(\text{s})$

Answer: D Diff: 3

4) Which of the following are strong acids? HI,  $\text{HNO}_3$ , HF, HBr

A) HF, HBr

C) HI, HF, HBr

E) HI,  $\text{HNO}_3$ , HBr

B) HI,  $\text{HNO}_3$ , HF, HBr

D)  $\text{HNO}_3$ , HF, HBr

Answer: E Diff: 2

5) A neutralization reaction between an acid and a metal hydroxide produces \_\_\_\_\_.

A) water and a salt

C) oxygen gas

E) ammonia

B) hydrogen gas

D) sodium hydroxide

Answer: A Diff: 2

6) Which of these metals is the least easily oxidized? Na, Au, Fe, Ca, Ag

A) Na

B) Au

C) Fe

D) Ca

E) Ag

Answer: B Diff: 3

7) Based on the equations below, which metal is the most active?



A) Ni

B) Ag

C) Cu

D) Pb

E) N

Answer: A Diff: 3

8) What is the concentration (M) of KCl in a solution made by mixing 25.0 mL of 0.100 M KCl with 50.0 mL of 0.100 M KCl?

A) 0.100

B) 0.0500

C) 0.0333

D) 0.0250

E) 125

Answer: A Diff: 3

# Modern Chemistry Chapter 8 Review Answer Key

**Léo Pomar**



## **Modern Chemistry Chapter 8 Review Answer Key:**

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**Foundations of College Chemistry, International Adaptation** Morris Hein, Susan Arena, Cary Willard, 2023-05-16 *Foundations of College Chemistry* 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students. It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence. The central focus is to make chemistry interesting and understandable and teach students the problem-solving skills they will need. This International Adaptation offers new and updated content with improved presentation of all course material. It builds on the strengths of previous editions, including clear explanations and step-by-step problem solving. The material emphasizes real-world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry. There is new and expanded coverage of polarizing power and polarizability, Fajans' rules, collision number, and mean free path, abnormal molecular masses, and van't Hoff factor and applications of radioactivity.

*Medicinal Chemistry* Norma K. Dunlap, Donna M. Hury, 2025-11-28 This book provides a concise introduction to the medicinal chemistry of therapeutically active compounds. Written in an accessible style, it offers a brief overview of the drug development process from discovery to approved drug. Emphasizing medicinal chemistry strategies, the content is organized by target as well as therapeutic area drug classes. It covers existing drugs and the discovery process that led to their development. Aimed at students and professionals interested in drug discovery, *Medicinal Chemistry* requires little prior knowledge, making it ideal for classroom and training seminar use. Written by experts in the field, this outstanding new edition explains medicinal chemistry on chemical mechanism terms, allowing for a more interesting and user-friendly text. Includes valuable insights toward the various pathways taken by the pharmaceutical industry in drug discoveries. Enhanced text includes questions raised and suggestions made from students in the author's medicinal chemistry class. Content

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