

## Lab 22A

### Investigating Chemical Equilibrium

**Purpose:** To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the the systems, which are exploring the structures through applying Le Chatelier's principle.

#### Proble:

1. Write a balanced equation for the reaction in part 1.  
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.  
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.  
$$\text{FeSCN}^{2+}(\text{aq}) + \text{H}^{+}(\text{aq}) \rightleftharpoons \text{FeSCNH}^{3+}(\text{aq})$$
4. In part 4, what is the purpose of adding potassium thiocyanate to the solution of cobalt ions in water?  
Cobalt ions have an orange color and form the hexa-aqua cobalt ion. The purpose of adding thiocyanate ions is to form the thiocyanate cobalt ion, which has a red color.
5. What is the purpose of adding distilled water in part 4?  
Water is added to dilute the solution, which causes the color to become lighter.
6. What is the purpose of the test tubes in part 4?  
To observe any color change in the solution as the equilibrium shifts.
7. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$$
8. Write any reactions that  $\text{Fe}^{3+}$  undergoes with  $\text{OH}^{-}$  ions.  
Cobalt ions  $\text{Co}^{2+}$  will react with hydroxide ions to form cobalt hydroxide  $\text{Co}(\text{OH})_2$  and water  $\text{H}_2\text{O}$  in the solution.
9. What is a precipitate?  
A precipitate is a solid that forms from a solution when two soluble ionic compounds react to form an insoluble solid.

# Investigating Chemical Equilibrium Lab 12a Answer

**Michael Rajnik**



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