

1-6, 8, 9, 11

## Hydrolysis of Salts and pH of Buffer Solutions

## Pre-lab Questions

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

1. Define Brønsted-Lowry acids and bases.

According to Brønsted-Lowry definitions, an acid is a proton donor and a base is a proton acceptor.

2

- Which of the following ions will react with water in a hydrolysis reaction:

$K^+$ ,  $Ba^{2+}$ ,  $Cu^{2+}$ ,  $Zn^{2+}$ ,  $F^-$ ,  $SO_3^{2-}$ , and  $Cl^-$ ?

$Cu^{2+}$ ,  $Zn^{2+}$ ,  $F^-$  and  $SO_3^{2-}$  will undergo hydrolysis.

3

- For those ions in question 2 that undergo hydrolysis, write net ionic equations for the hydrolysis reaction.

Example:  $Cu^{2+}(aq) + 2H_2O(l) \rightleftharpoons Cu(OH)_2(s) + 2H^+(aq)$ .

4

- The  $K_a$  for HCN is  $4.9 \times 10^{-10}$ . What is the value of  $K_b$  for  $CN^-$ ?

$$K_b = \frac{1.0 \times 10^{-14}}{4.9 \times 10^{-10}} = 2.0 \times 10^{-5}$$

5. What are the conjugate base and conjugate acid of  $H_2PO_4^-$ ?

The conjugate base is  $HPO_4^{2-}$  and the conjugate acid is  $H_3PO_4$ .

6

- From what acid and what base were the following salts made:  $K_2SO_4$ ,  $NH_4Cl$ , and  $Ca(NO_3)_2$ ?

For  $Ca^{2+}$  and  $SO_4^{2-}$  salt is made from  $Ca(OH)_2$  and  $H_2SO_4$ .

7

- Define the term *salt*.

# Ph Properties Of Buffer Solutions Answer Key Pre Lab

**R Barnett**



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,1970 *Industrial Laboratory*, 1958 **British Abstracts**, 1938 **Journal of Imaging Technology**, 1985 *British Ceramic Abstracts*, 1981 **The Planter and Sugar Manufacturer**, 1926 **Louisiana Planter and Sugar Manufacturer**, 1926 **Food Science and Technology Abstracts**, 1974 Monthly References from world literature of books about 1000 journals and patents from 18 selected countries Classified arrangement according to 18 sections such as milk and dairy products eggs and egg products and food microbiology Author subject indexes

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