

## Acid-Base Titration Lab

### Introduction

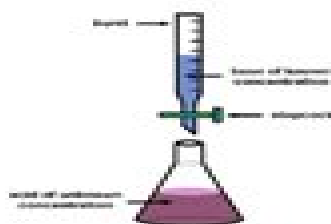
In chemistry laboratory, it is sometimes necessary to experimentally determine the concentration of an unknown acid or base solution. A procedure for making this kind of determination is called an **acid-base titration**. In this laboratory process, a solution of known concentration, called the standard solution, is carefully added to a solution of unknown concentration until the mixture becomes neutral. The neutral point of the solution is recognized by an indicator's color change. If the unknown solution is acidic, then the standard solution will be basic. The opposite would be true if the unknown solution was basic.

We know that the mixing of equal amounts of acid and base ions will create neutral water. At the molecular level, this reaction can be illustrated with the following equation.



This equation states that one mole of hydrogen ions (acid) will neutralize one mole of hydroxide ions (base). Since we can exactly measure the moles of the standard solution, we can assume that the moles of the solution of unknown concentration will be the **same** at the neutral point. This is called the **end-point** of the titration. Using the equation  $M_aV_a = M_bV_b$ , we can use the experimental data from the titration to find the unknown concentration.

**Objective:** In this experiment the unknown solution will be HCl(aq) and the standard solution will be the base sodium hydroxide. You will know the concentration of the base and the volume of the acid and base used. With this information you can use the titration formula to calculate the concentration of the acid. The diagram below shows the set up.



Swirl the flask continuously until one drop of titrant causes a color change to remain.  
A white sheet of paper can be placed under the flask to help you recognize the color change at the end-point.

### Pre Lab

1. Define the following words:

Titration- \_\_\_\_\_  
Endpoint- \_\_\_\_\_  
Neutralization- \_\_\_\_\_

2. Write the neutralization reaction for HCl reacting with NaOH.
3. What is the pH of the solution at the end point of the titration?

### Procedure

1. The flask is filled with 10 mL of unknown concentration of HCl. **Record the volume of acid on your data form**
2. Phenolphthalein indicator is added to the flask.
3. **Record on your data form** the initial volume of base in the buret
4. You will now start adding base (0.25M NaOH) from the buret into the flask to neutralize the acid. The flask is gently swirled as the base is added. A pink color should appear as the base is added, but it will disappear as the flask is swirled. The end-point of the titration will be reached when 1 drop of base makes the solution turn and stay pink.
5. **Record the final volume of base** from the buret that was needed to neutralize the acid.

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