

Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

Chemical Bonding:

- **Four types of bonds:**
 - *Nonpolar Covalent Bond* (ΔEN (difference in electronegativity) = < 0.3)
 - EQUAL SHARING of electrons between the 2 bonding nonmetal atoms
 - *Polar Covalent Bond* ($\Delta EN = 0.4-1.7$)
 - UNEQUAL SHARING of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
 - *Ionic Bond* ($\Delta EN = > 1.7$)
 - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely GIVING UP/TRANSFERRING ELECTRONS to the atom with the greater EN value
 - Ionic bond results due to the attraction between the POSITIVE cation (formed by losing electrons) and the NEGATIVE anion (formed by gaining electrons)
 - *Metallic Bond*
 - Bonding between two metallic atoms which occurs due to the attraction between a “sea of electrons” and the nuclei of bonded metal atoms
- * Must memorize/know the difference in electronegativity ranges for each type of bond
 - ** Note: you should still know the definition of electronegativity (EN)
- **Properties of the Four Types of Chemical Bonds**
 - See the table at the bottom of Packet pg.5
 - You MUST be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness
- **Why & How do chemical bonds form?**
 - Why → because all atoms want to be at their LOWEST POTENTIAL ENERGY; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
 - How → ATTRACTIVE and REPULSIVE forces
 - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
 - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repel one another (because they have the same charge)
 - **Bond Length**
 - ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
 - **Bond Energy**
 - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

Chapter 6 Study Guide Chemistry

Jianjun Gao



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Table of Contents Chapter 6 Study Guide Chemistry

1. Understanding the eBook Chapter 6 Study Guide Chemistry
 - The Rise of Digital Reading Chapter 6 Study Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Study Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 6 Study Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Study Guide Chemistry
 - Personalized Recommendations
 - Chapter 6 Study Guide Chemistry User Reviews and Ratings
 - Chapter 6 Study Guide Chemistry and Bestseller Lists
5. Accessing Chapter 6 Study Guide Chemistry Free and Paid eBooks
 - Chapter 6 Study Guide Chemistry Public Domain eBooks
 - Chapter 6 Study Guide Chemistry eBook Subscription Services
 - Chapter 6 Study Guide Chemistry Budget-Friendly Options

6. Navigating Chapter 6 Study Guide Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 6 Study Guide Chemistry Compatibility with Devices
 - Chapter 6 Study Guide Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 6 Study Guide Chemistry
 - Highlighting and Note-Taking Chapter 6 Study Guide Chemistry
 - Interactive Elements Chapter 6 Study Guide Chemistry
8. Staying Engaged with Chapter 6 Study Guide Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 6 Study Guide Chemistry
9. Balancing eBooks and Physical Books Chapter 6 Study Guide Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 6 Study Guide Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 6 Study Guide Chemistry
 - Setting Reading Goals Chapter 6 Study Guide Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 6 Study Guide Chemistry
 - Fact-Checking eBook Content of Chapter 6 Study Guide Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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