

## Chapter 6 Thermochemistry

### Homework Problems

- The enthalpy change for the fermentation (without oxygen,  $O_2$  gas) reaction of glucose to alcohol and carbon dioxide is  $-67$  kJ.
  - Write a chemical equation for the fermentation reaction and show how energy is involved in the reaction.
  - Is the reaction exothermic or endothermic?
  - Is energy, in the form of heat, absorbed or evolved as the reaction occurs?
- The specific heat capacity of graphite is  $0.71$  J/ $^{\circ}C$ g. Calculate the energy required to:
  - Raise the temperature of  $1.0$  mol graphite by  $1.0^{\circ}C$ .
  - Raise the temperature of  $850$  g graphite by  $150^{\circ}C$ .
  - Raise the temperature of  $75$  kg graphite from  $294$  K to  $348$  K.
- In a coffee-cup calorimeter,  $100.0$  mL of  $1.0$  M NaOH and  $100.0$  mL of  $1.0$  M HCl are mixed. Both solutions were originally at  $24.6^{\circ}C$ . After the reaction, the final temperature is  $31.3^{\circ}C$ . Assume that all the solutions have a density of  $1.0$  g/mL and a specific heat capacity of  $4.18$  J/g $^{\circ}C$ . Calculate the enthalpy change for the neutralization of HCl by NaOH. Assume that no heat is lost to the surroundings or to the calorimeter.
- Camphor ( $C_{15}H_{26}O$ ) has an energy of combustion of  $-5903.6$  kJ/mol. A sample of camphor with a mass of  $0.1204$  g is burned in a bomb calorimeter, the temperature increases by  $2.28^{\circ}C$ . Calculate the heat capacity of the calorimeter.  
A  $0.1964$  g sample of quinone ( $C_6H_4O_2$ ) is burned in a bomb calorimeter that has a heat capacity of  $1.56$  kJ/ $^{\circ}C$ . The temperature of the calorimeter increases by  $3.2^{\circ}C$ . Calculate the energy of combustion of quinone per gram and per mole.

# Chapter 6 Thermochemistry Homework Problems

**Michael Rajnik**



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### **Table of Contents Chapter 6 Thermochemistry Homework Problems**

1. Understanding the eBook Chapter 6 Thermochemistry Homework Problems
  - The Rise of Digital Reading Chapter 6 Thermochemistry Homework Problems
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Thermochemistry Homework Problems
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 6 Thermochemistry Homework Problems
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Thermochemistry Homework Problems
  - Personalized Recommendations
  - Chapter 6 Thermochemistry Homework Problems User Reviews and Ratings
  - Chapter 6 Thermochemistry Homework Problems and Bestseller Lists
5. Accessing Chapter 6 Thermochemistry Homework Problems Free and Paid eBooks
  - Chapter 6 Thermochemistry Homework Problems Public Domain eBooks
  - Chapter 6 Thermochemistry Homework Problems eBook Subscription Services

- Chapter 6 Thermochemistry Homework Problems Budget-Friendly Options
- 6. Navigating Chapter 6 Thermochemistry Homework Problems eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 6 Thermochemistry Homework Problems Compatibility with Devices
  - Chapter 6 Thermochemistry Homework Problems Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 6 Thermochemistry Homework Problems
  - Highlighting and Note-Taking Chapter 6 Thermochemistry Homework Problems
  - Interactive Elements Chapter 6 Thermochemistry Homework Problems
- 8. Staying Engaged with Chapter 6 Thermochemistry Homework Problems
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 6 Thermochemistry Homework Problems
- 9. Balancing eBooks and Physical Books Chapter 6 Thermochemistry Homework Problems
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 6 Thermochemistry Homework Problems
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 6 Thermochemistry Homework Problems
  - Setting Reading Goals Chapter 6 Thermochemistry Homework Problems
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 6 Thermochemistry Homework Problems
  - Fact-Checking eBook Content of Chapter 6 Thermochemistry Homework Problems
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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