



COVALENT BONDING

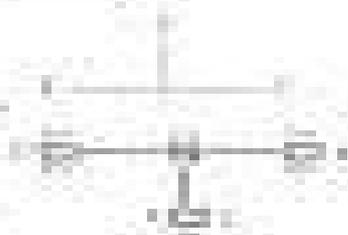
Practice Problems

Do your answers to the following problems.

SECTION 5.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - C_2H_6
 - H_2SO_4
- Write the formula for the following molecular compounds.
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide
- Write the name for each of the following molecular compounds.
 - N_2O_5
 - CO_2
 - CO
 - NO
 - NO_2
 - NO_3
- Write the name for each of the following molecular compounds.
 - C_2H_6
 - C_2H_4
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
 - $\text{C}_2\text{H}_6\text{O}_2$
 - $\text{C}_2\text{H}_4\text{O}_2$
- Write the formula for each of the following molecular compounds.
 - methane
 - ethane
 - propane
 - butane
 - pentane
 - hexane
 - heptane
 - octane
 - nonane
 - decane

SECTION 5.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride. $\text{H}-\text{F}$
- Show covalent bonding in methane. 
- Show covalent bonding in carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Show covalent bonding in ethane. $\text{H}-\text{C}-\text{C}-\text{H}$
- Show covalent bonding in ethyne. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for each of the following.
 - nitrogen
 - oxygen
 - hydrogen peroxide
 - water
 - hydrogen cyanide
 - carbon dioxide
 - carbon monoxide
 - nitrogen dioxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide
 - carbon disulfide
 - carbon tetrachloride
 - carbon tetrabromide
 - carbon tetrachloride
 - carbon tetrabromide
 - carbon tetrachloride
 - carbon tetrabromide



SECTION 5.3 BONDING: THE ORBITALS

- Which orbital has the highest energy in a carbon atom, $2s$ or the degenerate $2p$ orbitals?
- Which orbital has the highest energy in a nitrogen atom, $2s$ or the degenerate $2p$ orbitals?

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Table of Contents Chapter 8 Covalent Bonding Practice Problems Answers Bing

1. Understanding the eBook Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - The Rise of Digital Reading Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Personalized Recommendations
 - Chapter 8 Covalent Bonding Practice Problems Answers Bing User Reviews and Ratings
 - Chapter 8 Covalent Bonding Practice Problems Answers Bing and Bestseller Lists
5. Accessing Chapter 8 Covalent Bonding Practice Problems Answers Bing Free and Paid eBooks
 - Chapter 8 Covalent Bonding Practice Problems Answers Bing Public Domain eBooks
 - Chapter 8 Covalent Bonding Practice Problems Answers Bing eBook Subscription Services

- Chapter 8 Covalent Bonding Practice Problems Answers Bing Budget-Friendly Options
- 6. Navigating Chapter 8 Covalent Bonding Practice Problems Answers Bing eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 8 Covalent Bonding Practice Problems Answers Bing Compatibility with Devices
 - Chapter 8 Covalent Bonding Practice Problems Answers Bing Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Highlighting and Note-Taking Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Interactive Elements Chapter 8 Covalent Bonding Practice Problems Answers Bing
- 8. Staying Engaged with Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 8 Covalent Bonding Practice Problems Answers Bing
- 9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Practice Problems Answers Bing
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Setting Reading Goals Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Fact-Checking eBook Content of Chapter 8 Covalent Bonding Practice Problems Answers Bing
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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