

## VSA QUESTIONS (1 - MARK QUESTIONS)

1. Define the term 'rate of reaction'.
2. Mention the units of rate of reaction.
3. Express the rate of reaction in terms of  $\text{Br}^- (\text{aq})$  as reactant and  $\text{Br}_2 (\text{aq})$  as product for the reaction :



4. For a chemical reaction represented by  $\text{R} \rightarrow \text{P}$  the rate of reaction is denoted by

$$\frac{-\Delta[\text{R}]}{\Delta t} \text{ or } \frac{+\Delta[\text{P}]}{\Delta t}$$

Why a positive sign (+) is placed before  $\frac{\Delta[\text{P}]}{\Delta t}$  and negative sign (-)

before  $\frac{\Delta[\text{R}]}{\Delta t}$  ?

5. Express the rate of reaction in terms of disappearance of hydrogen and appearance of ammonia in the given reaction.



6. Why rate of reaction does not remain constant throughout?
7. Write the unit of first order rate constant of a gaseous reaction if the partial pressure of gaseous reactant is given in bar.
8. For a zero order reaction :

$\text{R} \rightarrow \text{P}$ , the change in concentration of reactant w.r.t. time is shown by following graph.

# Chemical Kinetics Questions And Answers

**Liying Dong**



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## **Table of Contents Chemical Kinetics Questions And Answers**

1. Understanding the eBook Chemical Kinetics Questions And Answers
  - The Rise of Digital Reading Chemical Kinetics Questions And Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Kinetics Questions And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Kinetics Questions And Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Kinetics Questions And Answers
  - Personalized Recommendations
  - Chemical Kinetics Questions And Answers User Reviews and Ratings
  - Chemical Kinetics Questions And Answers and Bestseller Lists
5. Accessing Chemical Kinetics Questions And Answers Free and Paid eBooks
  - Chemical Kinetics Questions And Answers Public Domain eBooks
  - Chemical Kinetics Questions And Answers eBook Subscription Services
  - Chemical Kinetics Questions And Answers Budget-Friendly Options
6. Navigating Chemical Kinetics Questions And Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Chemical Kinetics Questions And Answers Compatibility with Devices
  - Chemical Kinetics Questions And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemical Kinetics Questions And Answers
    - Highlighting and Note-Taking Chemical Kinetics Questions And Answers
    - Interactive Elements Chemical Kinetics Questions And Answers
  8. Staying Engaged with Chemical Kinetics Questions And Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemical Kinetics Questions And Answers
  9. Balancing eBooks and Physical Books Chemical Kinetics Questions And Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemical Kinetics Questions And Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemical Kinetics Questions And Answers
    - Setting Reading Goals Chemical Kinetics Questions And Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemical Kinetics Questions And Answers
    - Fact-Checking eBook Content of Chemical Kinetics Questions And Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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