

Chapter 4 – Atomic Structure

Short Answer: Answer the following questions in complete sentences using science vocabulary. Make sure that your answers are complete and detailed.

1. A sample of calcium contains calcium-40, calcium-44, calcium-42, calcium-48, calcium-43, and calcium-46 atoms. Explain why these atoms have different mass numbers but must have the same atomic number.

The atoms have different mass numbers because they have a different number of neutrons. The mass number is equaled to the number of protons plus neutrons.

The atoms must have the same atomic number because the atomic number is equaled to the number of protons. The number of protons is unique for each element.

2. Complete the following table. Circle the name of the element(s) that are isotopes. Complete Bohr's Model diagrams for the element with eleven electrons and the element with an atomic number of 9.

Element Name	Element Symbol	Atomic Number	Mass Number	Protons	Neutrons	Electrons
Scandium	Sc	21	45	21	24	21
Fluorine	F	9	19	9	10	9
Sodium	Na	11	24	11	13	11
Answers Vary					82	

Chapter 4 Atomic Structure Answers

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Chapter 4 Atomic Structure Answers:

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Solutions Manual Joseph Topich,2000-10 *On the Application of the Quantum Theory to Atomic Structure* Niels Bohr,1924 *Study Guide [to Accompany] General Chemistry* Dixie Goss,2002 *MCAT Comprehensive Review* ,2004 **Instructor's Guide** Alan Sherman,Leonard Russikoff,Sharon Sherman,1976 *Quantum Theory of Atomic Structure* John Clarke Slater,1960 *Earth Science* Samuel N. Namowitz,Nancy E. Spaulding,1989

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