

# ATOMIC STRUCTURE

Particle	Symbol	Charge	Relative charge	Mass	Relative mass w.r.t H	$\frac{e}{m}$ ratio
Electron	$e^-$	$-1.6 \times 10^{-19} \text{ C}$	-1	$9.1 \times 10^{-28} \text{ gram}$	$\frac{1}{1837}$	1837
Proton	$p^+$	$+1.6 \times 10^{-19} \text{ C}$	+1	$1.672 \times 10^{-24} \text{ g}$	1	1
Neutron	$n^0$	Zero	Zero	$1.672 \times 10^{-24} \text{ g}$	1	0
$\alpha$ -Particle	${}_2\text{He}^+$	$2 \times 1.6 \times 10^{-19} \text{ C}$	+2	$4 \times 1.672 \times 10^{-24} \text{ g}$	4	$\frac{1}{2}$

## # Thomson's atomic model

(Plum pudding / Watermelon model)

- +ve charge  $\rightarrow$  uniformly distributed
- -ve charge  $\rightarrow$  embedded in cloud of the charge.
- Limitation  $\rightarrow$  doesn't explain stability

## # Rutherford's experiment

(Discovery of nucleus)

Atom  $\begin{cases} \rightarrow \text{Nucleus (Proton)} \\ \rightarrow \text{Extranuclear part (e}^-) \end{cases}$

$$R_{\text{nucleus}} = K(A)^{1/3}$$

$$\left(\frac{R_{\text{nuc}}\right)_1 = \left(\frac{A_1}{A_2}\right)^{1/3}$$

$$V_{\text{nucleus}} = \frac{4}{3} \pi R^3$$

$$\frac{V_{\text{nuc}1}}{V_{\text{nuc}2}} = \frac{A_1}{A_2}$$

$$V_{\text{atom}} \approx 10^{15} \times V_{\text{nucleus}}$$

## # Term associated with element

$Z$   $\begin{cases} \nearrow A \rightarrow \text{Mass number (P+n)} \\ \searrow \text{Atomic number (P)} \end{cases}$

- **Isotope**  $\rightarrow Z \rightarrow$  same  $A \rightarrow$  different
- **Isobar**  $\rightarrow Z \rightarrow$  different  $A \rightarrow$  same
- **Isotone**  $\rightarrow A - Z = n \rightarrow$  same
- **Isotectonic**  $\rightarrow$  Same no. of  $e^-$  ( $\text{NO}^+$ ,  $\text{CO}$ )
- **Isosteric**  $\rightarrow$  Same no. of atom and  $e^-$  ( $\text{CO}_2$ ,  $\text{N}_2\text{O}$ )
- **Isotophers**  $\rightarrow (n-p) \text{ or } (A-2Z) \rightarrow$  same
- **Paramagnetic**  $\rightarrow$  Having non zero unpaired  $e^-$
- **Diamagnetic**  $\rightarrow$  Having zero unpaired  $e^-$

## # Dual nature of light

Wave nature

• Wave length ( $\lambda$ )  
distance between 2 crest or trough or length of one wave.

• Wave no.  $\left[ \nu = \frac{1}{\lambda} \right]$

• frequency  $\left[ \nu = \frac{c}{\lambda} \right]$

$c \rightarrow 3 \times 10^8 \text{ m/s}$

$\nu \propto \frac{1}{\lambda} \rightarrow \left[ \frac{\nu_1}{\nu_2} = \frac{\lambda_2}{\lambda_1} \right]$

Particle nature

• photon  $\rightarrow$  quanta of light.

$$E = h\nu$$

$h \rightarrow$  plank's constant  
 $h = 6.6 \times 10^{-34} \text{ joule sec}$

$$E = n \times E_{\text{min}}$$

$\downarrow$   
integral multiple

$$E = h\nu = \frac{hc}{\lambda}$$

$$\frac{E_1}{E_2} = \frac{\nu_1}{\nu_2} \quad \frac{E_1}{E_2} = \frac{\lambda_2}{\lambda_1}$$

## # Photoelectric effect

- example of particle nature.
- Surface phenomena
- Single photon release single  $e^-$

$$h\nu = \omega + KE \quad \omega \rightarrow \text{work function}$$

$$h\nu = h\nu_0 + \frac{1}{2} m_e v_e^2 \quad \nu \rightarrow \text{incident frequency of photon}$$

$$\frac{hc}{\lambda} = \frac{hc}{\lambda_0} + \frac{1}{2} m_e v_e^2 \quad \nu_0 \rightarrow \text{Threshold frequency of metal.}$$

#  $\nu > \nu_0 \rightarrow e^-$  emitted with KE  
( $\lambda < \lambda_0$ )

#  $\nu = \nu_0 \rightarrow e^-$  emitted without KE  
( $\lambda = \lambda_0$ )

#  $\nu < \nu_0 \rightarrow e^-$  not emitted.  
( $\lambda > \lambda_0$ )

# As Chemistry Revision Notes Unit 1 Atomic Structure

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