

## 11-2 Practice Problems

1. Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:



$$\frac{0.38\text{g Li}_3\text{N}}{1} \times \frac{1\text{mol Li}_3\text{N}}{35\text{g Li}_3\text{N}} \times \frac{3\text{mol LiOH}}{1\text{mol Li}_3\text{N}} \times \frac{24\text{g LiOH}}{1\text{mol LiOH}} = \boxed{0.78\text{g LiOH}}$$

2. What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?



$$\frac{0.29\text{g NaI}}{1} \times \frac{1\text{mol}}{150} \times \frac{2\text{mol NaCl}}{2\text{mol NaI}} \times \frac{58\text{g}}{1\text{mol}} = \boxed{0.112\text{g NaCl}}$$

3. Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:



$$\frac{0.85\text{g C}_4\text{H}_{10}}{1} \times \frac{1\text{mol}}{58\text{g}} \times \frac{8\text{mol CO}_2}{2\text{mol C}_4\text{H}_{10}} \times \frac{44\text{g}}{1\text{mol}} = \boxed{2.58\text{g CO}_2}$$

4. Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:



$$\frac{0.46\text{g Sb}_2\text{O}_3}{1} \times \frac{1\text{mol}}{291.52\text{g}} \times \frac{2\text{mol Sb}}{1\text{mol Sb}_2\text{O}_3} \times \frac{121.76\text{g}}{1\text{mol}} = \boxed{0.38\text{g Sb}}$$

5. What mass of hydrogen peroxide ( $\text{H}_2\text{O}_2$ ) must decompose to produce 0.77 g of water?



$$\frac{0.77\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{2\text{H}_2\text{O}_2}{2\text{H}_2\text{O}} \times \frac{34\text{g}}{1\text{mol}} = \boxed{1.45\text{g H}_2\text{O}_2}$$

6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?



$$\frac{0.69\text{g CO}_2}{1} \times \frac{1\text{mol}}{44\text{g}} \times \frac{2\text{mol CO}}{2\text{mol CO}_2} \times \frac{28\text{g}}{1\text{mol}} = \boxed{0.44\text{g CO}}$$

7. Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:



$$\frac{0.73\text{g Ni}(\text{NO}_3)_2}{1} \times \frac{1\text{mol}}{185\text{g}} \times \frac{2\text{mol NaNO}_3}{1\text{mol Ni}(\text{NO}_3)_2} \times \frac{85\text{g}}{1\text{mol}} = \boxed{0.68\text{g NaNO}_3}$$

8. Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:



$$\frac{0.64\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{1\text{mol Ca}(\text{OH})_2}{2\text{mol H}_2\text{O}} \times \frac{74\text{g}}{1\text{mol}} = \boxed{1.32\text{g Ca}(\text{OH})_2}$$

9. How many grams of ozone ( $\text{O}_3$ ) must decompose to produce 0.87 g of oxygen?



$$\frac{0.87\text{g O}_2}{1} \times \frac{1\text{mol}}{32\text{g O}_2} \times \frac{2\text{mol}}{3\text{mol}} \times \frac{48\text{g O}_3}{1\text{mol}} = \boxed{0.87\text{g O}_3}$$

10. Find the mass of sugar ( $\text{C}_6\text{H}_{12}\text{O}_6$ ) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation:



11. How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP? The balanced equation is  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$ .

12. Find the mass of benzene ( $\text{C}_6\text{H}_6$ ) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:



# Prentice Hall Chemistry Workbook Answers Chapter 10

**Jeanette Adkins**



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**Our Changing Social Order** Ruth Wood Gavian, Robert Rienow, 1964 **Paleoclimatology** Colin P.

Summerhayes, 2020-06-10 Life on our planet depends upon having a climate that changes within narrow limits not too hot for the oceans to boil away nor too cold for the planet to freeze over Over the past billion years Earth's average temperature has stayed close to 14-15°C oscillating between warm greenhouse states and cold icehouse states We live with variation but a variation with limits Paleoclimatology is the science of understanding and explaining those variations those limits and the forces that control them Without that understanding we will not be able to foresee future change accurately as our population grows Our impact on the planet is now equal to a geological force such that many geologists now see us as living in a new geological era the Anthropocene Paleoclimatology describes Earth's passage through the greenhouse and icehouse worlds of the past 800 million years including the glaciations of Snowball Earth in a world that was then free of land plants It describes the operation of the Earth's thermostat which keeps the planet fit for life and its control by interactions between greenhouse gases land plants chemical weathering continental motions volcanic activity orbital change and solar variability It explains how we arrived at our current understanding of the climate system by reviewing the contributions of scientists since the mid 1700s showing how their ideas were modified as science progressed And it includes reflections based on the author's involvement in palaeoclimatic research The book will transform debate and set the agenda for the next generation of thought about future climate change It will be an invaluable course reference for undergraduate and postgraduate students in geology climatology oceanography and the history of science A real tour de force An outstanding summary not only of the science and what needs to be done but also the challenges that are a consequence of psychological and cultural baggage that threatens not only the survival of our own species but the many others we are eliminating as well Peter Barrett Emeritus Professor of Geology Antarctic Research Centre Victoria University of Wellington New Zealand What a remarkable and wonderful synthesis it will be a wonderful source of paleoclimate information and insights Christopher R Scotese Professor Department of Earth and Planetary Sciences Northwestern University Evanston IL USA **Mining and Metallurgy**, 1925

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In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Prentice Hall Chemistry Workbook Answers Chapter 10**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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