

Chapter 10: Chemical analysis

Knowledge organiser

Pure and impure

In chemistry, a **pure** substance contains a single element or compound that is not mixed with any other substance.

Pure substances melt and boil at specific temperatures.

An **impure** substance contains more than one type of element or compound in a **mixture**.

Impure substances melt and boil at a range of temperatures.

Formulations

Formulations are examples of mixtures. They have many different components (substances that make them up) in very specific proportions (amounts compared to each other).

Scientists spend a lot of time trying to get the right components in the right proportions to make the most useful product.

Formulations include fuels, cleaning agents, paints, alloys, fertilisers, and foods.



Testing gases

Common gases can be identified using the follow tests:

Gas	What you do	What you observe if gas is present
hydrogen	hold a lighted splint near the gas	hear a squeaky pop
oxygen	hold a glowing splint near the gas	splint re-lights
carbon dioxide	bubble the gas through limewater	the limewater turns milky (cloudy white)
chlorine	hold a piece of damp litmus near the gas	litmus/blee litmus white

Chromatography

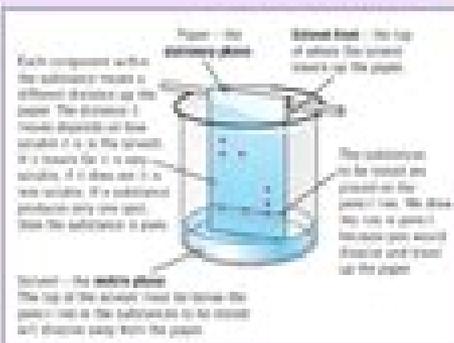
Chromatography is a method to separate different components in a mixture. It's set up as shown here, with a piece of paper in a beaker containing a small amount of solvent.

The **R_f value** is a ratio of how far up the paper a certain spot moves compared to how far the solvent has travelled.

$$R_f = \frac{\text{distance moved by substance}}{\text{distance moved by solvent}}$$

R will always be a number between 0 and 1.

The R_f value depends on the solvent and the temperature, and different substances will have different R_f values. The R_f values for particular solvents can be used to identify a substance.



Separating mixtures

Mixtures can be separated by:

- Filtration** - separates insoluble solids from a liquid
- crystallisation** - evaporates a solvent (liquid) leaving the solute (solid)
- simple distillation** - separates solvent from a solution as long as the solvent has a lower boiling point than the solute
- Fractional distillation** - separates two or more substances from a liquid mixture



Key terms

Make sure you can write a definition for these key terms.

chromatography formulation impure multiple phase precipitate pure R_f value
 solvent solvent front stationary phase

Chapter 10 Study Notes Docx Flancoore

Thomas Griffiths



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