

Chapter 12

SECTION 12.1

12.1.1 Boyle's Law: Pressure-Volume Relationship for an Ideal Gas

1. Boyle's Law states that pressure of a fixed mass of gas is inversely proportional to its volume. Show that the following are equivalent:

(a) Pressure is inversely proportional to volume at constant temperature.

(b) Pressure is directly proportional to the reciprocal of volume at constant temperature.

(c) The product of pressure and volume is constant at constant temperature.

(d) Boyle's Law is a special case of the general gas law.

2. Two sealed flasks, A and B, contain the same amount of gas at the same temperature. The volume of flask A is twice that of flask B. The pressure in flask A is P_A and the pressure in flask B is P_B . What is the ratio of P_A to P_B ? (Hint: Use Boyle's Law.)

3. A gas is compressed from a volume of 10 L to a volume of 5 L. The pressure is 1 atm. What is the new pressure?

4. A gas is compressed from a volume of 10 L to a volume of 5 L. The pressure is 1 atm. What is the new pressure?

5. A gas is compressed from a volume of 10 L to a volume of 5 L. The pressure is 1 atm. What is the new pressure?

6. A gas is compressed from a volume of 10 L to a volume of 5 L. The pressure is 1 atm. What is the new pressure?

7. A gas is compressed from a volume of 10 L to a volume of 5 L. The pressure is 1 atm. What is the new pressure?

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