

14

THE BEHAVIOR OF GASES

Chapter Goals

After studying this chapter, you should be able to:

- 1. Explain Boyle's, Charles's, and Gay-Lussac's laws, the ideal gas law, and the combined gas law.
- 2. Calculate molecular weights using the ideal gas law, and determine molecular weights using density.
- 3. Calculate the partial pressure, mole fraction, and volume of a gas in a mixture of gases.
- 4. Calculate the molar mass and percent composition of a gas, and determine the partial pressure of a gas in a mixture of gases.
- 5. Determine the mass of a gas from the ideal gas law, and determine the volume of a gas.

After studying this chapter, you should be able to:

- 1. Calculate the mass of a gas from the ideal gas law, and determine the volume of a gas.

Problems

- 1. A certain gas occupies a volume of 1.00 L at 1.00 atm and 0°C. What volume will it occupy at 2.00 atm and 0°C? What volume will it occupy at 1.00 atm and 100°C? What volume will it occupy at 2.00 atm and 100°C?

Problems

- 1. A certain gas occupies a volume of 1.00 L at 1.00 atm and 0°C. What volume will it occupy at 2.00 atm and 0°C? What volume will it occupy at 1.00 atm and 100°C? What volume will it occupy at 2.00 atm and 100°C?

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Fred M. Dewey



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