

CHAPTER 15 REVIEW

Chemical Equilibrium

Teacher Notes and Answers

Chapter 15 Introduction 1

Question 1

- a. $\frac{[CO_2]}{[CO]}$
- b. $\frac{[CO_2]}{[CO]}$

- c. $\frac{[CO_2]}{[CO]}$
- d. $\frac{[CO_2]}{[CO]}$

e. $\frac{[CO_2]}{[CO]}$ (change of)

f. $\frac{[CO_2]}{[CO]}$

g. $\frac{[CO_2]}{[CO]}$ (change of)

- h. $\frac{[CO_2]}{[CO]}$
- i. $\frac{[CO_2]}{[CO]}$

- 10. a. The value of the equilibrium constant, K_c , is affected by temperature and the concentrations of the products and the reactants.

- b. K_c
- c. K_c

Introduction 2

Question 1

- a. K_c
- b. K_c
- c. K_c

- d. K_c
- e. K_c
- f. K_c
- g. K_c

- 10. These graphs are not suitable representations of equilibrium constant K_c versus temperature. The graphs show K_c increasing with temperature and K_c decreasing with T .

- 11. a. The value of K_c depends on the Gibbs free energy change for the reaction and the standard enthalpy change.

Calculating equilibrium constants from the standard enthalpy change and the standard Gibbs free energy change is not given.

Calculating equilibrium constants from the standard enthalpy change and the standard Gibbs free energy change is not given.

- b. The equilibrium constant is the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.

- c. The equilibrium constant is the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.

15.1 Equilibrium

Question 1

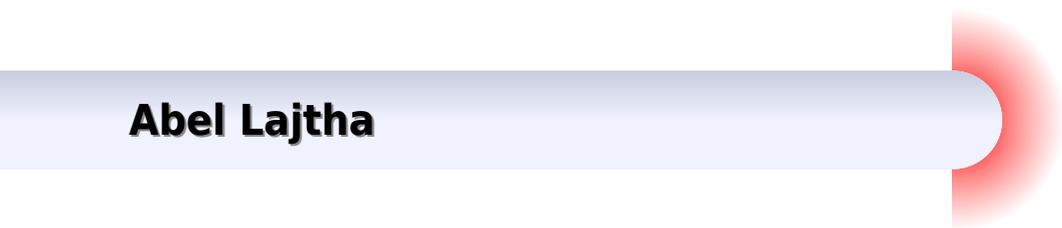
- a. K_c
- b. K_c
- c. K_c
- d. K_c
- e. K_c
- f. K_c
- g. K_c
- h. K_c
- i. K_c
- j. K_c
- k. K_c
- l. K_c
- m. K_c
- n. K_c
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- r. K_c
- s. K_c
- t. K_c
- u. K_c
- v. K_c
- w. K_c
- x. K_c
- y. K_c
- z. K_c

Question 2

- a. K_c
- b. K_c
- c. K_c
- d. K_c
- e. K_c
- f. K_c
- g. K_c
- h. K_c
- i. K_c
- j. K_c
- k. K_c
- l. K_c
- m. K_c
- n. K_c
- o. K_c
- p. K_c
- q. K_c
- r. K_c
- s. K_c
- t. K_c
- u. K_c
- v. K_c
- w. K_c
- x. K_c
- y. K_c
- z. K_c

Chapter 18 Review Chemical Equilibrium Section 3 Answers

Abel Lajtha



Chapter 18 Review Chemical Equilibrium Section 3 Answers:

Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **Alterations of Chemical Equilibrium in the Nervous System** Abel Lajtha, 2013-11-21 It has been recognized for more than a thousand years that the function of the brain like the function of the other organs of the body is determined by its physical chemical and biological properties Evidence that even its highest functions could be explained by these properties was gathered only in recent years however these findings which clearly have to be confirmed by a great deal of further experimental evidence indicate that most if not all of the functions of the brain are based on its bio chemical and biophysical mechanisms This at first hearing may sound rather simple but the ability to understand learning emotion perhaps even creativity on biological terms may well be the most important scientific discovery of all time Few pieces of knowledge can influence our future health and well being to the degree that understanding of mental mechanisms will It has been clearly shown in many ways in the previous volumes of this Handbook that from the biochemical or neurochemical point of view the brain is one of the most active organs The brain seems stable and in some respects permanent this is evidence not of inactivity but of carefully controlled homeostasis of dynamic rather than static equilibrium with most components undergoing metabolic alterations

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