

25. The number of σ bonds in $N \equiv N$ is
- 1
 - 2
 - 3
 - 4
26. The elements in an ionic compound are held together by
- electrostatic forces of attraction.
 - van der Waals forces
 - the spin of paired electrons.
 - the formation of hybrid orbitals.
 - an electron pair.
27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode
- at which oxidation occurs.
 - which attracts cations.
 - at which electrons are supplied to the solution.
 - at which reduction occurs.
28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?
- 47
 - 62
 - 93
 - 186
29. In a galvanic (voltaic) cell in which the reaction is $Cd + Cu^{2+} \rightarrow Cu + Cd^{2+}$ and the ions are at unit concentration (activity), the cell potential is
- $$Cd \rightarrow Cd^{2+} + 2e^- \quad 0.4021 \text{ V}$$
- $$Cu \rightarrow Cu^{2+} + 2e^- \quad -0.344 \text{ V}$$
- 0.1383 V
 - 0.4021 V
 - 0.344 V
 - 0.7461 V
 - 0.3677 V
30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?
- $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$
 - $H_2(g) + Cl_2(g) \rightleftharpoons 2HCl(g)$
 - $2NO(g) + O_2(g) \rightleftharpoons 2NO_2(g)$
 - $COCl_2(g) \rightleftharpoons CO(g) + Cl_2(g)$

Standard Potentials	E°
$Mg \rightarrow Mg^{2+} + 2e^-$	2.37V
$Al \rightarrow Al^{3+} + 3e^-$	1.66V
$Zn \rightarrow Zn^{2+} + 2e^-$	0.76V
$Fe \rightarrow Fe^{2+} + 2e^-$	0.44V
$Cu \rightarrow Cu^{2+} + 2e^-$	0.34V
$Ag \rightarrow Ag^+ + e^-$	-0.80V

31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.
- Mg and Ag
 - Mg and Fe
 - Zn and Cu
 - Al and Ag
 - Mg and Al
32. $E = E^\circ - 0.059/n \log Q$ (Nernst equation)
 $[H^+] = 1.0 \text{ M}$ initially, $P_{O_2} = 1.0 \text{ atm}$
- $$4e^- + O_2(g) + 4H^+(aq) \rightleftharpoons 2H_2O(l) \quad E^\circ = 1.23 \text{ V}$$
- Based on the information above, which statement is correct?
- $n = 1$, since one mole of oxygen is being considered.
 - Addition of base should result in an E value, which is less than 1.23 V.
 - E is independent of the pH of the solution.
 - $Q = \frac{[H_2O]^2}{[O_2][H^+]^4}$
33. The equilibrium constant for the gaseous reaction $C + D \rightleftharpoons E + 2F$ is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will
- proceed at equal rates in both directions.
 - proceed more rapidly to form E and F.
 - proceed more rapidly to form C and D.
 - not occur in either direction.
- | Compound | ΔG_f° , kJ/mol |
|-----------|-----------------------------|
| $H_2O(l)$ | -237 |
| $H_2O(g)$ | -229 |
34. At 298 K the equilibrium constant for $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(l)$
- is larger than the K_{eq} for $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(g)$
 - will have a value of 1.0 at equilibrium.
 - cannot be computed since data on O_2 and

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Preparing for Your ACS Examination in General Chemistry - the Official Guide LUCY T. EUBANKS, 2018-10-15 This guide is separated into first term and second term general chemistry material Each section contains 8 chapters of material that also aligns to most general chemistry textbooks for a seamless addition to study materials for students Each chapter is designed with an introductory section of the material including common representations and where to find this material in a textbook The second section provides worked examples of typical multiple choice questions including how the correct answer is determined as well as how the incorrect answers were determined Also included for each study problem is a listing of the corresponding practice questions that use that concept The final section is a series of practice problems to test the concepts collectively The key is provided on a separate page for all study and practice problems

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McMurry, DonnaJean Fredeen, Robert Pribush, 1998-01-30

Teaching Innovation in University Education: Case

Studies and Main Practices Saura, Jose Ramon, 2022-06-17 In the last decade the development of new technologies has made innovation a fundamental pillar of education Teaching innovation includes the evolution of both teaching and learning models to drive improvements in educational methodologies Teaching innovation is a pioneer in the understanding and comprehension of the different teaching methodologies and models developed in the academic area Teaching innovation is a process that seeks validation in the academic and teaching communities at universities in order to promote the improvement and its practices and uses in the future characterized by digital development and data based methods Teaching Innovation in University Education Case Studies and Main Practices features the major practices and case studies of teaching innovation developed in recent years at universities It is a source on study cases focused on teaching innovation methodologies as well as on the identification of new technologies that will help the development of initiatives and practices focused on teaching innovation at higher education institutions Covering topics such as didactic strategics service learning and technology based gamification this premier reference source is an indispensable resource for pre service teachers lecturers students faculty administrators libraries entrepreneurs researchers and academicians

Subject Guide to Books in Print, 2001 ACS

Organic Chemistry Sterling Test Prep, 2022-10-14 ASC Organic Chemistry bestseller Thousands of students use Sterling Test Prep study aids to achieve high test scores High yield practice questions and detailed explanations for topics tested on ACS Organic Chemistry examination This book provides high yield practice questions covering organic chemistry topics Chemistry instructors with years of teaching experience prepared these questions by analyzing the test content and developing practice material that builds your knowledge and skills crucial for success on the ACS Our test preparation experts structured the content to match the current test requirements The detailed explanations describe why an answer is correct and more important for your learning why another attractive choice is wrong They provide step by step solutions and teach the important details of organic chemistry mechanisms and reactions needed to answer ACS exam questions Read the explanations carefully to understand how they apply to the question and learn important organic chemistry principles and the relationships between them Scoring well on ACS Organic Chemistry exam is a challenging task This book helps you develop

and apply knowledge to quickly choose the correct answer on the test Solving targeted practice questions builds your understanding of fundamental general chemistry concepts and is a more effective strategy than merely memorizing terms With this practice material you will significantly improve your test score

Advances in Teaching Physical Chemistry Mark David Ellison, 2008 This book brings together the latest perspectives and ideas on teaching modern physical chemistry It includes perspectives from experienced and well known physical chemists a thorough review of the education literature pertaining to physical chemistry a thorough review of advances in undergraduate laboratory experiments from the past decade in depth descriptions of using computers to aid student learning and innovative ideas for teaching the fundamentals of physical chemistry This book will provide valuable insight and information to all teachers of physical chemistry

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Visualization in Science Education John K. Gilbert, 2005-07-05 Visualization meaning both the perception of an object that is seen or touched and the mental imagery that is the product of that perception is believed to be a major strategy in all thought It is particularly important in science which seeks causal explanations for phenomena in the world as experienced Visualization must therefore play a major role in science education This book addresses key issues concerning visualization in the teaching and learning of science at any level in educational systems Visualization in Science Education draws on the insights from cognitive psychology science and education by experts from Australia Israel Slovenia UK and USA It unites these with the practice of science education particularly the ever increasing use of computer managed modelling packages especially in chemistry The first section explores the significance and intellectual standing of visualization The second section shows how the skills of visualization have been developed practically in science education This is followed by accounts of how the educational value of visualization has been integrated into university courses in physics genomics and geology The fourth section documents experimental work on the classroom assessment of visualization An endpiece summarises some of the research and development needed if the contribution of this set of universal skills is to be fully exploited at all levels and in all science subjects

The ETS Test Collection Catalog Educational Testing Service. Test Collection, 1993 The major source of information on the availability of standardized tests Wilson Library Bulletin Covers commercially available standardized tests and hard to locate research instruments

Guide to Educational Credit by Examination Douglas R. Whitney, Andrew G. Malizio, 1987

Making it relevant Peter Nentwig, David Waddington, 2006 Teaching in context has become an accepted and often welcomed way of teaching science in both primary and secondary schools The conference organised by IPN and the University of York Science Education Group Context based science curricula drew on the experience of over 40 science educators and 10 projects The book is arranged in four parts Part A consists of two papers one on situated learning and the other on implementation of new curricula Part B contains descriptions of five major curricula in different countries why they were introduced how they were developed and implemented and evaluation results Part C gives descriptions of three projects

that are of smaller scale and their materials are used as interventions in other more conventional curricula There is also a contribution on some fundamental research where modules of work are written to examine how best to design context based curricula Finally Part D consist of two chapters one summarising some of the findings that came out of the chapters in the three earlier parts and the second looks at the future Griswold Chem Im 6/E Nebergall,1980 Resources in Education,1988

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