

Exercise 2: Creating an Activity Series

Table 2: Redox Reactions of Copper, Lead, and Zinc.

Redox Metal	Redox Ox	Solution	Immediate Observations	30-Minute Observations
Cu	H ₂ O	Pb(NO ₃) ₂	No noticeable changes	None
	H ₂ O	Zn(NO ₃) ₂	No noticeable changes	None
Pb	H ₂ O	CuSO ₄	Blue solution, a rust color appeared on the surface	Blue color faded and appeared to be rusty
	H ₂ O	Zn(NO ₃) ₂	None	None
Zn	H ₂ O	CuSO ₄	Appeared to be dark, rusty, and fuzzy appearance, blue solution	Appeared to be rusted, was a dark colored almost black, less blue
	H ₂ O	Pb(NO ₃) ₂	Dark almost black, clear solution	Appeared to be fuzzy, dark, and grey colored, clear solution

Table 3: Potential Redox Reactions and Chemical Equations.

Metal and Aqueous Solution	Reaction Occurred?	Chemical Equation
Cu + Pb(NO ₃) ₂	No	No reaction
Cu + Zn(NO ₃) ₂	No	No reaction
Pb + CuSO ₄	Yes	$Pb + CuSO_4 \rightarrow Cu + PbSO_4$
Pb + Zn(NO ₃) ₂	No	$Pb + Zn(NO_3)_2 \rightarrow Zn + Pb(NO_3)_2$
Zn + CuSO ₄	Yes	$Zn + CuSO_4 \rightarrow Cu + ZnSO_4$
Zn + Pb(NO ₃) ₂	Yes	$Zn + Pb(NO_3)_2 \rightarrow Pb + Zn(NO_3)_2$

Activity Series Lab Report Answers

Lingjun Ying



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