

## Section 18.4 Neutralization

In your textbook, read about neutralization and titration.

For each item in Column A, write the letter of the matching item in Column B.

### Column A

- \_\_\_\_\_ 1. A chemical dye that changes color based on the pH of a solution
- \_\_\_\_\_ 2. A method for using a neutralization reaction to determine the concentration of a solution
- \_\_\_\_\_ 3. A reaction in which an acid and a base react to produce a salt and water
- \_\_\_\_\_ 4. A solution of known concentration
- \_\_\_\_\_ 5. An ionic product of an acid-base reaction
- \_\_\_\_\_ 6. The point in a titration in which an indicator changes color
- \_\_\_\_\_ 7. The stoichiometric point of a titration

### Column B

- a. acid-base indicator
- b. end point
- c. equivalence point
- d. neutralization
- e. salt
- f. standard solution
- g. titration

Complete the following table, indicating the formula and name of the salt formed by a neutralization reaction between the listed acid and base.

Acid	Base	Salt formula	Salt name
8. HCl	KOH	KCl	potassium chloride
9. $H_2SO_4$	$Mg(OH)_2$		
10. $H_3PO_4$	NaOH		
11. $HNO_3$	$Fe(OH)_3$		
12. $H_3PO_4$	$Ca(OH)_2$		

In the space at the left, write 1 through 4 to show the correct sequence of the steps in performing a titration using a pH meter. Then, write 5 through 8 to sequence the steps used to calculate the concentration of the unknown solution.

### Sequence of Steps

- \_\_\_\_\_ 13. Continue adding the standard solution to the solution of unknown concentration until the equivalence point is reached.
- \_\_\_\_\_ 14. Fill a buret with the standard solution.
- \_\_\_\_\_ 15. Start adding the standard solution slowly, with mixing, to the solution of unknown concentration, reading the pH at regular intervals.

# Chapter 9 Review Stoichiometry Section 1 Answer Key

**M Lipman**



## Chapter 9 Review Stoichiometry Section 1 Answer Key:

**Modern Chemistry** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001      **Chemistry** Oxtoby, 1994  
*Holt Chemistry* Ralph Thomas Myers, 2004      **Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry** Iris Stovall, Donald J. DeCoste, 2000      *Foundations of College Chemistry* Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years Foundations of College Chemistry Alternate 14th Edition has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis      McGraw-Hill's SAT Subject Test: Chemistry, 2ed Thomas Evangelist, 2009-02-01 We want to help you score high on the SAT Chemistry test We ve put all of our proven expertise into McGraw Hill s SAT Subject Test Chemistry to make sure you re fully prepared for this difficult exam With this book you ll get essential skill building techniques and strategies created by a leading high school chemistry teacher You ll also get 5 full length practice tests hundreds of sample questions and all the facts about the current exam With McGraw Hill s SAT Subject Test Chemistry we ll guide you step by step through your preparation program and give you the tools you need to succeed 5 full length sample tests with complete explanations for every question 40 top test items to remember on exam day A step by step review of all topics covered on the exam Teacher recommended tips and strategies to help you raise your score      **Study Guide for General Chemistry and College Chemistry, Eigth Editions by Holtzclaw and Robinson** Norman E. Griswold, Henry Fuller Holtzclaw, 1988      **McGraw-Hill's SAT Subject Test Chemistry, 3rd Edition** Thomas Evangelist, 2011-02-03 Expert guidance on the Chemistry exam Many colleges and universities require you to take one or more SAT II Subject Tests to demonstrate your mastery of specific high school subjects McGraw Hill s SAT Subject Test Chemistry is written by experts in the field and gives you the guidance you need perform at your best This book includes 4 full length sample tests updated for the latest test format 40 top tips to remember on test day Glossary of tested chemistry terms and formulas Tips and strategies from one of the most popular teachers at the renowned Brooklyn Technical High School Diagnostic test to pinpoint strengths and weaknesses Step by step review of all topics covered on the exam In depth coverage of the lab experiment questions that are a major test feature Charts tables and illustrations to simplify and reinforce learning Practice tests just like the real SAT Subject Test in Chemistry Test taking tips and strategies      *Math Review Toolkit for General Chemistry* Gary Long, 2006-05-30      Solving Problems in Analytical Chemistry Stephen Brewer, 1980 Basic material Useful statistics Stoichiometry Acid base chemistry General discussion Strong acids and bases Weak monoprotic acids Weak polyprotic acids The titration weak acids The titration of weak acids Weak bases The control of solubility Complex formation with a ligand

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