

## Chapter 9 Stoichiometry Chapter Review Answers

- Mole Ratios
  - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
  - They are formed from the coefficients found in a balanced chemical equation.
- Mole Ratios as conversion factors
  - $\frac{2 \text{ mol Ca}}{1 \text{ mol O}_2} \cdot \frac{1 \text{ mol O}_2}{2 \text{ mol Ca}} \cdot \frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} \cdot \frac{2 \text{ mol Ca}}{2 \text{ mol CaO}} \cdot \frac{1 \text{ mol O}_2}{2 \text{ mol CaO}} \cdot \frac{2 \text{ mol CaO}}{1 \text{ mol O}_2}$ . Remember the ratios can be reduced as well!!!
  - $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} \cdot \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol H}_2} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol MgF}_2} \cdot \frac{1 \text{ mol MgF}_2}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol H}_2} \cdot \frac{1 \text{ mol H}_2}{2 \text{ mol HF}} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol H}_2} \cdot \frac{1 \text{ mol H}_2}{1 \text{ mol MgF}_2}$
- $\frac{105.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \cdot \frac{1 \text{ mol Na}_2\text{CO}_3}{105.99 \text{ g Na}_2\text{CO}_3} \cdot \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \cdot \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} \cdot \frac{40.00 \text{ g NaOH}}{1 \text{ mol NaOH}} \cdot \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \cdot \frac{100.09 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} \cdot \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3}$
- Molar Mass
  - Definition: the mass of one mole of a substance
  - Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen  $\rightarrow$  Water
  - 5.0 mol H<sub>2</sub>
  - 2.5 mol O<sub>2</sub>
- Ethane combustion reaction
  - 15.8 mol O<sub>2</sub>
  - 9.00 mol CO<sub>2</sub> & 13.5 mol H<sub>2</sub>O
- 575 g Na & 886 g Cl<sub>2</sub>
- Iron production reaction
  - 75.1 mol CO
  - 50.1 mol Fe & 75.1 mol CO<sub>2</sub>
- 87,390 g CO or  $8.739 \times 10^4$  g CO & 12,610 g H<sub>2</sub> or  $1.261 \times 10^4$  g H<sub>2</sub>
- Lightning reaction of N<sub>2</sub> & O<sub>2</sub>
  - 1,100 g NO<sub>2</sub> or  $1.10 \times 10^3$  g NO<sub>2</sub>
  720. g NO or  $7.20 \times 10^2$  g NO
- CO<sub>2</sub> removal in spacecraft
  - 42.04 mol NaOH
  - 21.02 mol Na<sub>2</sub>CO<sub>3</sub> & 21.02 mol H<sub>2</sub>O
- Double replacement AgNO<sub>3</sub> & NaBr
  - 463 g NaBr
  - 845 g AgBr
- Soda-acid fire extinguisher
  - 3.058 mol NaHCO<sub>3</sub>
  - 3.058 mol CO<sub>2</sub>, 1.529 mol Na<sub>2</sub>SO<sub>4</sub>, 3.058 mol H<sub>2</sub>O
- Sulfuric acid-sodium hydroxide neutralization reaction
  - H<sub>2</sub>SO<sub>4</sub> + 2 NaOH  $\rightarrow$  Na<sub>2</sub>SO<sub>4</sub> + 2 H<sub>2</sub>O
  - 37 g H<sub>2</sub>SO<sub>4</sub>
  - 53 g Na<sub>2</sub>SO<sub>4</sub> & 14 g H<sub>2</sub>O
- Copper & silver nitrate reaction
  - 0.0104 mol Cu(NO<sub>3</sub>)<sub>2</sub>
  - 0.0104 mol Cu & 0.0209 mol AgNO<sub>3</sub>
- Aspirin production
  - 13.5 kg C<sub>9</sub>H<sub>8</sub>O<sub>4</sub>
  - 7.66 kg C<sub>7</sub>H<sub>6</sub>O<sub>3</sub>
  - 4.29 L HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>
- Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant **limits** the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

# Chapter 9 Review Stoichiometry Section 2 Answers

**AW Rasmussen**



## Chapter 9 Review Stoichiometry Section 2 Answers:

*Modern Chemistry* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 MCAT Comprehensive Review 1999, 1998-11 Complete Review of ALL the science you need to know Full Length Practice Test with complete explanations for each question 2 Full Length Verbal Reasoning Tests with detailed explanations and strategies **CliffsNotes AP Chemistry 2021 Exam** Angela Woodward Spangenberg, 2020-09-29 CliffsNotes AP Chemistry 2021 Exam gives you exactly what you need to score a 5 on the exam concise chapter reviews on every AP Chemistry subject in depth laboratory investigations and full length model practice exams to prepare you for the May 2021 exam Revised to even better reflect the new AP Chemistry exam this test prep guide includes updated content tailored to the May 2021 exam Features of the guide focus on what AP Chemistry test takers need to score high on the exam Reviews of all subject areas In depth coverage of the all important laboratory investigations Two full length model practice AP Chemistry exams Every review chapter includes review questions and answers to pinpoint problem areas Holt Chemistry Ralph Thomas Myers, 2004 *MCAT Comprehensive Review*, 2004 MCAT Comprehensive Review, 2004 Kaplan, Inc, 2003 The bestselling MCAT prep guide is a must for pre med students preparing for the demanding Medical College Admissions Test MCAT The CD features one full length practice test with detailed explanations plus effective tips and strategies *Basic Concepts of Chemistry, Study Guide* Leo J. Malone, 2000 The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems **AP Chemistry Premium, 2022-2023: Comprehensive Review with 6 Practice Tests + an Online Timed Test Option** Neil D. Jespersen, Pamela Kerrigan, 2021-07-06 Be prepared for exam day with Barron's Trusted content from AP experts Barron's AP Chemistry Premium 2022 2023 includes in depth content review and online practice It's the only book you'll need to be prepared for exam day Written by Experienced Educators Learn from Barron's all content is written and reviewed by AP experts Build your understanding with comprehensive review tailored to the most recent exam Get a leg up with tips strategies and study advice for exam day it's like having a trusted tutor by your side Be Confident on Exam Day Sharpen your test taking skills with 6 full length practice tests 3 in the book and 3 more online Strengthen your knowledge with in depth review covering all Units on the AP Chemistry Exam Reinforce your learning with practice questions at the end of each chapter Interactive Online Practice Continue your practice with 3 full length practice tests on Barron's Online Learning Hub Simulate the exam experience with a timed test option Deepen your understanding with detailed answer explanations and expert advice Gain

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