

20 Points

Name: Phyllis A. Smith Lab Section: 200

Suppose that you had the correct equation for the following combustion reaction. What highest boiling point compound?

$\text{C}_2\text{H}_6\text{O}$

C_2H_6

C_2H_4

C_2H_2

$\text{C}_2\text{H}_6\text{O}$
Ethanol

Diethyl ether

C_2H_4

C_2H_2
Acetylene

1. If you have the correct equation for the combustion of $\text{C}_2\text{H}_6\text{O}$, what is the highest boiling point compound in the reaction mixture?



2. If you were in the mixture and couldn't see the correct equation for $\text{C}_2\text{H}_6\text{O}$, what would be the highest boiling point compound in the mixture?



100 g H_2O

100 g H_2O

100 g H_2O

100 g H_2O

Chem 110 Lab Questions And Answers

J Spring



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Laboratory Instruction in the Field of Inorganic Chemistry Victor Herbert Noll,1930 **Fundamentals of Chemistry: Laboratory Studies** Frank Brescia,2012-12-02 Fundamentals of Chemistry Laboratory Studies Third Edition is a manual that provides instruction on techniques of chemical laboratory operations Each experiment is discussed in terms of the major objective the experimental approach to the objective the measurements or observations to be made and the calculation and interpretation of results Topics covered include manipulation weights and measures molecular weight acids and bases gravimetric and volumetric stoichiometry and thermochemistry This book is comprised of 43 chapters divided into 14 sections and begins by presenting general information on metric and other units common laboratory equipment and chemical laboratory methods The first chapter introduces the reader to the Bunsen burner and the principles of glass working followed by a discussion on mass and volume measurements including the determination of density The following chapters focus on states of matter molecular weight stoichiometry and intermolecular forces Preparations and syntheses are also considered along with chemical equilibrium and electrochemistry The final section is devoted to qualitative analysis particularly of cations and anions This monograph is intended primarily for students of chemistry **General Chemistry ... Laboratory Manual and Note Book** American School (Lansing, Ill.),1905 *Examination Questions in Biology, Botany, Chemistry, Drawing, Geography, Physics, Zoölogy* College Entrance Examination Board,1920 Laboratory Experiments in General Chemistry John Chandler,1981 *Laboratory Exercises for a Brief Course in Chemistry* Lyman Churchill Newell,1927 *Laboratory Experiments for Brown and LeMay, Chemistry, the Central Science* John Henry Nelson,1985

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