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September 28, 2021

**Title:** What Is Chemical Reaction?

**Introduction:** The goal of this experiment is to determine chemical properties of hydrochloric acid and copper chloride; identify the evidence that we can gather to find out which chemical change that occurred, whether in nature or laboratory; if the law of conservation of mass applies to a sample chemical reaction in one of these series. A chemical change is defined as a variable composition and properties of substance. There are many types of observable changes such as formation of solid precipitate upon mixing two solutions; release of gas bubbles that are not due to a physical change; a color change that does not result from dilution or color mixing; or a temperature change that is not caused by the external heating or cooling. For instance, a nail can become tarnished when it gets old as new substances begin to take the place of the old substances of the nail. Still, it is physically a nail but the "formula" differentiates. The nail now has been oxidized. If you are a cook, you are now a food scientist. Chemical changes don't have to occur in the lab all the time, they can actually be right there in your kitchen. Think of a simple breakfast recipe like scrambled eggs. Cooking an egg uses a chemical reaction because the heat polymerizes proteins in egg white, while the hydrogen and sulfur in the yolk can react to form hydrogen sulfide gas. Same thing happens with cooked meats, and cake.

**Objective:** to examine the chemical properties of hydrochloric acid and copper chloride, to identify the types of reactions that they undergo, and to determine the law of conservation of mass applies to a sample chemical reaction.

**Hypothesis:** as acid is a hazardous chemical, the samples might experience corroding or discolor as soaking in the hydrochloric solution. The samples may not react as dramatically to copper chloride as hydrochloric because it does not have an acid name.

**Result:**

Data Table A: Reactions of Hydrochloric Acid

Initial appearance of HCl Solutions: transparent, and clear liquid

Initial temperature of HCl solutions: 74° F

Reaction Well	Reagents	Observations
1	HCl + NaOH	After 1 minute: 72° F Litmus paper turns pink.
2	HCl + NaHCO <sub>3</sub>	Bubble frizzing, temperature decreases
3	HCl + AgNO <sub>3</sub>	Foam occurs in the beginning, then the solution

# Chemical Reaction Lab Report Grade 10

**Michael Rajnik**



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