

## Answers to Chapter 3 Study Questions

1. a)  $3(12.0) + 8(1.01) + 3(16.0) = 92.1 \text{ g/mole}$       b)  $92.1 \text{ g}$       c)  $6.02 \times 10^{23}$  molecules  
 d)  $0.217 \text{ moles} \times \frac{92.1 \text{ g}}{1 \text{ mole}} = 20.0 \text{ g}$
2. a) 4 atoms (one N + 3 H)  
 b)  $1 \text{ mole } \text{NH}_3 \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 2.41 \times 10^{24} \text{ atoms}$   
 c)  $3.40 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 4.82 \times 10^{23} \text{ atoms}$
3. Molar mass of  $\text{NaNO}_2 = 23.0 + 14.0 + 2(16.0) = 69.0 \text{ g/mole}$   
 $\% \text{ Na} = 23.0/69.0 = 33.3\% \text{ Na}$ ;  $\% \text{ N} = 14.0/69.0 = 20.3\% \text{ N}$ ;  $\% \text{ O} = 2(16.0)/69.0 = 46.4\% \text{ O}$   
**33.3% Na, 20.3% N and 46.4% O.**
4. a) In 100 g of this compound, there are 40.7 g C, 5.1 g H, and 54.2 g O  
 $40.7 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 3.39 \text{ moles C}$        $3.39/3.39 = 1 \times 2 = 2$   
 $5.1 \text{ g H} \times \frac{1 \text{ mole H}}{1.0 \text{ g H}} = 5.1 \text{ moles H}$        $5.1/3.39 = 1.5 \times 2 = 3$   
 $54.2 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 3.39 \text{ moles O}$        $3.39/3.39 = 1 \times 2 = 2$   
 empirical formula =  **$\text{C}_2\text{H}_3\text{O}_2$**
- b) Molar mass of  $\text{C}_2\text{H}_3\text{O}_2 = 2(12.0) + 3(1.0) + 2(16.0) = 59.0 \text{ g/mole}$   
 $118/59.0 = 2 \rightarrow$  molecular formula =  **$\text{C}_4\text{H}_6\text{O}_4$**
5. a) In 25.0 g of compound, there are 7.20 g Mg, 3.55 g C and  $25.0 - (7.20 + 3.55) = 14.25 \text{ g O}$ .  
 $7.20 \text{ g Mg} \times \frac{1 \text{ mole Mg}}{24.3 \text{ g Mg}} = 0.296 \text{ moles Mg}$        $0.296/0.296 = 1$   
 $3.55 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 0.296 \text{ moles C}$        $0.296/0.296 = 1$   
 $14.25 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 0.891 \text{ moles O}$        $0.891/0.296 = 3$   
 formula =  **$\text{MgCO}_3$**
- b)  $\% \text{ Mg} = 7.20/25.0 = 28.8\% \text{ Mg}$ ;  $\% \text{ C} = 3.55/25.0 = 14.2\% \text{ C}$ ;  $\% \text{ O} = 14.25/25.0 = 57.0\% \text{ O}$ . (You should get the same result using molar mass and atomic masses.)
- c)  $13.9 \text{ g compound} \times \frac{28.8 \text{ g Mg}}{100 \text{ g compound}} = 4.00 \text{ g Mg}$

# Chemistry Chapter 3 Worksheet Answers

**SJ Ball**



**Chemistry Chapter 3 Worksheet Answers:**

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In some sort of used by screens and the ceaseless chatter of fast interaction, the melodic splendor and psychological symphony created by the written word usually diminish into the back ground, eclipsed by the constant noise and disturbances that permeate our lives. However, set within the pages of **Chemistry Chapter 3 Worksheet Answers** a wonderful fictional prize brimming with raw feelings, lies an immersive symphony waiting to be embraced. Crafted by an outstanding musician of language, that captivating masterpiece conducts readers on an emotional trip, skillfully unraveling the hidden songs and profound affect resonating within each carefully crafted phrase. Within the depths of the emotional analysis, we shall explore the book is central harmonies, analyze their enthralling writing design, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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goddess of agriculture in one legend her daughter persephone is carried off by hades the god of the underworld

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web demeter in greek religion a consort of zeus and the goddess of agriculture especially grain though rarely mentioned by homer and not an olympian deity she is probably an ancient goddess she is best remembered for her role in the story of persephone in which her lack of attention to the harvest causes a famine

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