

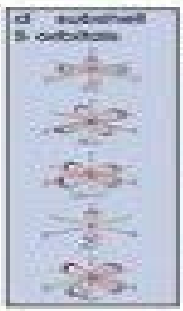
CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box:

- 1** — atomic number
- H** — atomic symbol
- 1.008** — atomic weight



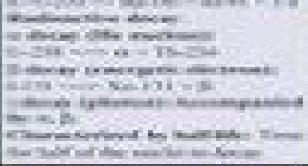
ATOMIC STRUCTURE

Atomic Number, Z , is of protons in the nucleus. For neutral atom, $Z = +$ of electrons. For ions of charge, $+Z$ (cations) or $-Z$ (anions). For ions of charge, $+Z$ (cations) or $-Z$ (anions). Mass Number, A , is $Z + N$, the number of nucleons in the nucleus. Isotopes atoms with the same Z , different N , the changed sign have a number of protons, may also differ in a number of neutrons. In practical work use the average Mass Number or Atomic Weight (weighted average of natural isotopes for a given element).

NUCLEAR CHEMISTRY

Nuclear reactions alter the number, particles present or chemical behavior of nuclei and have energy. Energy (E) is related to mass (m) using the Einstein equation, $E = mc^2$.
Nuclei are built from protons, p , and/or neutrons, n , and are symbolized by ${}^A_Z X$, or ${}^A_Z X$ (e.g., Deuterium, ${}^2_1\text{D}$ or ${}^2_1\text{H}$).

Types of processes:
 Transmutation: Two elements made from particle collisions.
 Fusion (small nuclei combine): ${}^2_1\text{H} + {}^3_1\text{H} \rightarrow {}^4_2\text{He} + {}^1_0\text{n}$
 Radioactive decay:
 α decay: ${}^A_Z X \rightarrow {}^{A-4}_{Z-2} Y + {}^4_2\text{He}$
 β^- decay: ${}^A_Z X \rightarrow {}^A_{Z+1} Y + e^- + \bar{\nu}_e$
 β^+ decay: ${}^A_Z X \rightarrow {}^A_{Z-1} Y + e^+ + \nu_e$
 Electron capture: ${}^A_Z X + e^- \rightarrow {}^A_{Z-1} Y + \nu_e$



ATOMIC QUANTUM NUMBERS AND ORBITALS

Quantum numbers describe the energy, size, and shape of an orbital. The principal quantum number, n , is the most important. It determines the energy level and the size of the orbital. The azimuthal quantum number, l , determines the shape of the orbital. The magnetic quantum number, m_l , determines the orientation of the orbital. The spin quantum number, m_s , determines the spin of the electron in the orbital.

Orbitals are labeled by their principal quantum number, n , and azimuthal quantum number, l . For example, a 2p orbital has $n=2$ and $l=1$. The number of orbitals in a subshell is given by $2l+1$. For example, there are 3 p orbitals ($l=1$) and 5 d orbitals ($l=2$).

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe all many-electron atoms and ions. A many-electron atom has many different energy levels and orbital properties. The main point is that orbitals have different energies, and the filling of levels is guided by the Aufbau Principle.


n	l	ml	orbitals
1	0	0	1 orbital
2	0	0	1 orbital
2	1	0, ±1	3 orbitals
3	0	0	1 orbital
3	1	0, ±1	3 orbitals
3	2	0, ±1, ±2	5 orbitals
4	0	0	1 orbital
4	1	0, ±1	3 orbitals
4	2	0, ±1, ±2	5 orbitals
4	3	0, ±1, ±2, ±3	7 orbitals

Atomic configuration: Shows orbital occupancy. Pauli Exclusion Principle: Each orbital has a unique set of quantum numbers, and can hold only two electrons, one with spin up, \uparrow , and one with spin down, \downarrow . Hund's Rule: Electrons fill orbitals singly, as a maximizes the number of unpaired electrons. Aufbau Principle: Orbitals are filled in order of increasing energy. The order of filling is given by the Aufbau Principle. The order of filling is given by the Aufbau Principle.

Large periodic table with color-coded groups and detailed element information.

Chemistry Quickstudy Reference Guides Academic

**Mark Jackson (PhD and Science
Professor)**



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