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so calculating enthalpy change of reaction in 2 $\Delta H_{\text{reaction}} = 6 \times \Delta H_{\text{f}}^{\circ}(\text{CH}_4) - 6 \times \Delta H_{\text{f}}^{\circ}(\text{C}_6\text{H}_6) + 6 \times \Delta H_{\text{f}}^{\circ}(\text{CO}_2) + 6 \times \Delta H_{\text{f}}^{\circ}(\text{H}_2\text{O})$

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