

TABLE A-11

Saturated refrigerant-134a—Temperature table

Temp., T °C	Sat. pres., P _{sat} kPa	Specific volume, m ³ /kg		Internal energy, kJ/kg			Enthalpy, kJ/kg			Entropy, kJ/kg · K		
		Sat. liquid, v _f	Sat. vapor, v _g	Sat. liquid, u _f	Evap., u _{fg}	Sat. vapor, u _g	Sat. liquid, h _f	Evap., h _{fg}	Sat. vapor, h _g	Sat. liquid, s _f	Evap., s _{fg}	Sat. vapor, s _g
-40	51.25	0.0007054	0.34081	-0.036	207.40	207.37	0.000	225.86	225.86	0.00000	0.96866	0.96866
-38	56.86	0.0007083	0.32732	2.475	206.04	208.51	2.535	224.61	227.12	0.01072	0.95511	0.96584
-36	62.95	0.0007112	0.29751	4.992	204.67	209.66	5.037	223.35	228.39	0.02138	0.94176	0.96315
-34	69.56	0.0007142	0.27090	7.517	203.29	210.81	7.566	222.09	229.65	0.03199	0.92859	0.96058
-32	76.71	0.0007172	0.24711	10.05	201.91	211.96	10.10	220.81	230.91	0.04253	0.91560	0.95813
-30	84.43	0.0007203	0.22580	12.59	200.52	213.11	12.65	219.52	232.17	0.05301	0.90278	0.95579
-28	92.76	0.0007234	0.20666	15.13	199.12	214.25	15.20	218.22	233.43	0.06344	0.89012	0.95356
-26	101.73	0.0007265	0.18946	17.69	197.72	215.40	17.76	216.92	234.68	0.07382	0.87762	0.95144
-24	111.37	0.0007297	0.17395	20.25	196.30	216.55	20.33	215.59	235.92	0.08414	0.86527	0.94941
-22	121.72	0.0007329	0.15995	22.82	194.88	217.70	22.91	214.26	237.17	0.09441	0.85307	0.94748
-20	132.82	0.0007362	0.14729	25.39	193.45	218.84	25.49	212.91	238.41	0.10463	0.84101	0.94564
-18	144.69	0.0007396	0.13583	27.98	192.01	219.98	28.09	211.55	239.64	0.11481	0.82908	0.94389
-16	157.38	0.0007430	0.12542	30.57	190.56	221.13	30.69	210.18	240.87	0.12493	0.81729	0.94222
-14	170.93	0.0007464	0.11597	33.17	189.09	222.27	33.30	208.79	242.09	0.13501	0.80561	0.94063
-12	185.37	0.0007499	0.10736	35.78	187.62	223.40	35.92	207.38	243.30	0.14504	0.79406	0.93911
-10	200.74	0.0007535	0.099516	38.40	186.14	224.54	38.55	205.96	244.51	0.15504	0.78263	0.93766
-8	217.08	0.0007571	0.092352	41.03	184.64	225.67	41.19	204.52	245.72	0.16498	0.77130	0.93629
-6	234.44	0.0007608	0.085802	43.66	183.13	226.80	43.84	203.07	246.91	0.17489	0.76008	0.93497
-4	252.85	0.0007646	0.079804	46.31	181.61	227.92	46.50	201.60	248.10	0.18476	0.74896	0.93372
-2	272.36	0.0007684	0.074304	48.96	180.08	229.04	49.17	200.11	249.28	0.19459	0.73794	0.93253
0	293.01	0.0007723	0.069255	51.63	178.53	230.16	51.86	198.60	250.45	0.20439	0.72701	0.93139
2	314.84	0.0007763	0.064612	54.30	176.97	231.27	54.55	197.07	251.61	0.21415	0.71616	0.93031
4	337.90	0.0007804	0.060338	56.99	175.39	232.38	57.25	195.51	252.77	0.22387	0.70540	0.92927
6	362.23	0.0007845	0.056398	59.68	173.80	233.48	59.97	193.94	253.91	0.23356	0.69471	0.92828
8	387.88	0.0007887	0.052762	62.39	172.19	234.58	62.69	192.35	255.04	0.24323	0.68410	0.92733
10	414.89	0.0007930	0.049403	65.10	170.56	235.67	65.43	190.73	256.16	0.25286	0.67356	0.92641
12	443.31	0.0007975	0.046295	67.83	168.92	236.75	68.18	189.09	257.27	0.26246	0.66308	0.92554
14	473.19	0.0008020	0.043417	70.57	167.26	237.83	70.95	187.42	258.37	0.27204	0.65266	0.92470
16	504.58	0.0008066	0.040748	73.32	165.58	238.90	73.73	185.73	259.46	0.28159	0.64230	0.92389
18	537.52	0.0008113	0.038271	76.08	163.88	239.95	76.52	184.01	260.53	0.29112	0.63198	0.92310

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An Inductive Approach to Engineering Thermodynamics George Sidebotham, 2022-03-28 This textbook provides an alternative inductive treatment of traditional Engineering Thermodynamics e.g energy and its transformations in engineering systems and introduces the notion of eXergy The book begins with energy methods developed in mechanics and transitions to thermodynamics by introducing both 1st and 2nd Laws of Thermodynamics immediately incorporating more advanced concepts using practical applications This methodology continues throughout the text wherein consideration of a specific example leads to general conclusions At the same time the author introduces eXergy also called Availability a measure of the potential of a substance to produce useful mechanical work in being brought from its current state to the conditions of the local environment The book facilitates students understanding with workshop problem statements and guided spreadsheet It is appropriate for a sophomore or junior level first course in thermodynamics and is restricted to simple compressible substances with no formal chemical reaction development Mechanical engineering applications are the primary target where several follow up courses would follow fluid mechanics heat transfer and a 2nd thermos course Civil or electrical engineering students could benefit from just this course and chemical engineering programs could develop chemically reacting and non ideal applications in follow up courses

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Property Tables Booklet to Accompany Thermodynamics Yunus A. Çengel,Michael A. Boles,2008 **Property Tables and Figures** Kenneth Wark,1999 **Property Tables and Figures for Engineering Thermodynamics** Iowa State University. Department of Mechanical Engineering,2000 **Forthcoming Books** Rose Army,1989-09 *Thermodynamics Anfd Property Tables* Yungus A. Cengel,1998-01-01 [Thermodynamic Tables, Bibliography, and Property File \(Sections VII, VIII, and IX\)](#) Harold L. Schick,1966 **Thermodynamic Property Tables and Worked Tutorial Problems in Thermodynamics** Kenneth Charles Brown,1998 [Thermodynamic Property Tables](#) Bruce Fegley,Harvard College Observatory, *Handbook of Thermodynamic Tables and Charts* K. Raznjevic,1976 This volume provides comprehensive treatment of basic data for engineering thermodynamics as well as practical and theoretical thermodynamics All numerical values are cited in both technical and SI units Contents are arranged according to aggregate states solids liquids vapors and gases And at the end of the book is a table converting units of measure from one system to another Pref Information on thermal properties melting points specific heat thermal conductivity etc Twenty one pages of conversion factors and equivalents Two large charts in inside back cover on water enthalpy vs entropy Indexed Published 1976 *Handbook of Thermodynamic Tables and Charts* Kuzman Raznjevic,1976 **Handbook of Thermodynamic Tables and Diagrams** Charles Edward Lucke,1915 **Multiparameter Equations of State** Roland Span,2010-12-15 As a basis for printed property charts and tables empirical multiparameter equations of state have been the most important source of accurate thermodynamic property data for more than 30 years now However due to increasing demands on the accuracy of thermodynamic property data in computerised calculations as well as the availability of appropriate software tools and the ever increasing computer power such formulations are nowadays becoming a valuable tool for everyday work This development has substantially increased the number of scientists engi neers and students who are working with empirical multiparameter equations of state and it continues to do so Nevertheless common knowledge on this kind of thermodynamic property models and on the ongoing progress in this scientific discipline is still very limited Multiparameter equations of state do not belong to the topics which are taught intensively in thermodynamic courses in engineering and

natural sciences and the books and articles where they are published mainly deal with the thermodynamic properties of certain substances rather than with the theoretical background of the used equations of state In contrast to this my concern mainly was to give a survey of the theoretical background of multiparameter equations of state both with regard to their application and their development

National Bureau of Standards,1947 Tables of Selected Values of Chemical Thermodynamic Properties United States
Project,1971 *International Thermodynamic Tables of the Fluid State* Thermodynamic Tables

Table of Contents Thermodynamics Property Tables Cengel 7th Edition

1. Understanding the eBook Thermodynamics Property Tables Cengel 7th Edition
 - The Rise of Digital Reading Thermodynamics Property Tables Cengel 7th Edition
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermodynamics Property Tables Cengel 7th Edition
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermodynamics Property Tables Cengel 7th Edition
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermodynamics Property Tables Cengel 7th Edition
 - Personalized Recommendations
 - Thermodynamics Property Tables Cengel 7th Edition User Reviews and Ratings
 - Thermodynamics Property Tables Cengel 7th Edition and Bestseller Lists
5. Accessing Thermodynamics Property Tables Cengel 7th Edition Free and Paid eBooks
 - Thermodynamics Property Tables Cengel 7th Edition Public Domain eBooks
 - Thermodynamics Property Tables Cengel 7th Edition eBook Subscription Services
 - Thermodynamics Property Tables Cengel 7th Edition Budget-Friendly Options
6. Navigating Thermodynamics Property Tables Cengel 7th Edition eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermodynamics Property Tables Cengel 7th Edition Compatibility with Devices
 - Thermodynamics Property Tables Cengel 7th Edition Enhanced eBook Features
7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Thermodynamics Property Tables Cengel 7th Edition
 - Highlighting and Note-Taking Thermodynamics Property Tables Cengel 7th Edition
 - Interactive Elements Thermodynamics Property Tables Cengel 7th Edition
8. Staying Engaged with Thermodynamics Property Tables Cengel 7th Edition
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermodynamics Property Tables Cengel 7th Edition
 9. Balancing eBooks and Physical Books Thermodynamics Property Tables Cengel 7th Edition
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermodynamics Property Tables Cengel 7th Edition
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Thermodynamics Property Tables Cengel 7th Edition
 - Setting Reading Goals Thermodynamics Property Tables Cengel 7th Edition
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Thermodynamics Property Tables Cengel 7th Edition
 - Fact-Checking eBook Content of Thermodynamics Property Tables Cengel 7th Edition
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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