

Advanced Chemistry useful handout #1

Metric conversion chart

10^9 g =	1 Gg	giga	
10^6 g =	1 Mg	mega	
10^3 g =	1 kg	kilo	$1000 \text{ g} = 1 \text{ kg}$
1 g =	1 g	BASE	
1 g =	10^1 dg	deci	$1 \text{ g} = 10 \text{ dg}$
1 g =	10^2 cg	centi	$1 \text{ g} = 100 \text{ cg}$
1 g =	10^3 mg	milli	$1 \text{ g} = 1000 \text{ mg}$
1 g =	10^6 μ g	micro	
1 g =	10^9 ng	nano	
1 g =	10^{12} pg	pico	

10^{10} Å (Angstroms) = 1 meter (atomic size)

Linear to Volume: When converting between linear and volume measurements, you need to *cube* the conversion factor! $10 \text{ cm} = 1 \text{ dm}$; so $(10 \text{ cm})^3 = (1 \text{ dm})^3$ thus $1000 \text{ cm}^3 = 1 \text{ dm}^3 = 1 \text{ Liter}$

Some Helpful Formulas:

Gas Laws: Conversions: STP for gases = 0°C and 1 atm ; (or 273 K , 760 mm Hg)

Combined : $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$ Diffusion: $\frac{\text{rate A}}{\text{rate B}} = \frac{\sqrt{FW B}}{\sqrt{FW A}}$ Molar volume of = $\frac{22.4 \text{ L}}{\text{mole}}$
 a gas at STP

Ideal Gas Law: $PV = nRT$ If you know P, V, T of a gas, you can calculate moles!

For $PV = nRT$: P in atm, V in L, n = moles, T in K, Ideal Gas Constant, R = $0.0821 \text{ L atm/mol K}$
 (use Partial Pressure of a gas in $PV = nRT$ to get mol of that particular gas)

Dalton's law of Partial Pressures: $P_{\text{total}} = P_a + P_b + P_c + \dots$ $P_{\text{dry gas}} = P_{\text{total}} - P_{\text{water vapor}}$

Partial pressure = $P_a = X_a P_{\text{total}}$ where mol fraction = $X_a = \frac{\text{mol a}}{\text{Total mol of gases}}$

Solutions: Molarity = $\frac{\text{moles solute}}{\text{L of solution}}$ or $M = \frac{\text{mol}}{\text{L}}$ mol = V M $V = \frac{\text{mol}}{M}$ (V must be in Liters)

Or : can use millimoles and milliliters: $M = \frac{\text{mmol}}{\text{mL}}$ mmol = mL M mL = $\frac{\text{mmol}}{M}$

Dilutions: $V_1 M_1 = V_2 M_2$ (can use mL for Volumes) Acid Base Titration: $V_a M_a (\#H) = V_b M_b (\#OH)$

molality = $m = \frac{\text{mol solute}}{\text{kg solvent}}$ mol fraction = $X_a = \frac{\text{mol a}}{\text{total mol}}$ Weight % = $\frac{\text{g solute}}{100 \text{ g solution}}$

Additional conversion factors and constants:

Energy:

$4.18 \text{ joules} = 1 \text{ calorie}$ $1 \text{ Joule} = \frac{1 \text{ kg m}^2}{\text{sec}^2}$

Volume: $1 \text{ cm}^3 = 1 \text{ mL}$ $1 \text{ dm}^3 = 1 \text{ L}$

$1 \text{ m}^3 = 10^3 \text{ L}$

Density = $\frac{\text{Mass}}{\text{Volume}}$; $\frac{1 \text{ gram}}{\text{mL}}$ (water only)

Avogadro's number: $N = \frac{6.02 \times 10^{23} \text{ particles}}{1 \text{ mole}}$

Kelvin = $^\circ\text{Celsius} + 273$

$1 \text{ atm} = 760 \text{ mm Hg}$

Advanced Chemistry

Guillaume Favre



Advanced Chemistry:

The Enigmatic Realm of **Advanced Chemistry**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Advanced Chemistry** a literary masterpiece penned with a renowned author, readers embark on a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting affect the hearts and minds of those who partake in its reading experience.

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