

# Chemistry of Life

## DOODLE NOTES

**chemistry of life**  
**BASIC CHEMISTRY REVIEW**

**atoms and life**  
All living things are made of atoms.  
An atom is the smallest part of an element. Many different types of atoms exist and they are arranged in the periodic table of elements. Atoms are made up of subatomic particles.

Label the structure of this oxygen atom:

• Note the charge and mass of each subatomic particle.

Charge	Relative mass	Relative charge
+	1	0
0	1	-1

• How many shells are in the first shell? 2  
• How many shells are in the outer shell? 2  
• What do we call electrons in the outer shell? Valence electrons

**elements**  
An element is a substance that is made of one type of atom only. They cannot be broken down into simpler substances.

**compounds**  
A compound is made of two or more types of atoms that are chemically bonded. They can be broken down chemically. There are both compounds and mixtures.

**mixtures**  
A mixture is a mix of two or more types of atoms that are NOT chemically bonded. They can be separated physically. There are solutions, suspensions, and emulsions.

**Examples:**  
Helium - He  
Sulfur - S  
Sodium Chloride - NaCl  
Oxide - O<sub>2</sub>

**ionic bonding**

Electron transfer from a metal to a non-metal.

**Example:** Sodium Chloride (NaCl)

Na: 1 valence electron, Cl: 7 valence electrons.

Na loses 1 electron, Cl gains 1 electron.

Na<sup>+</sup> Cl<sup>-</sup>

**functional groups**

These particular groups of atoms are attached to the carbon chain to give the properties that determine them. They are called functional groups.

**Alcohol**  
-OH

**Phosphate**  
-PO<sub>4</sub>

**Acids**  
-COOH



MULTIPLE VERSIONS & EXAMPLE

# Biology The Chemistry Of Life Answer Key

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