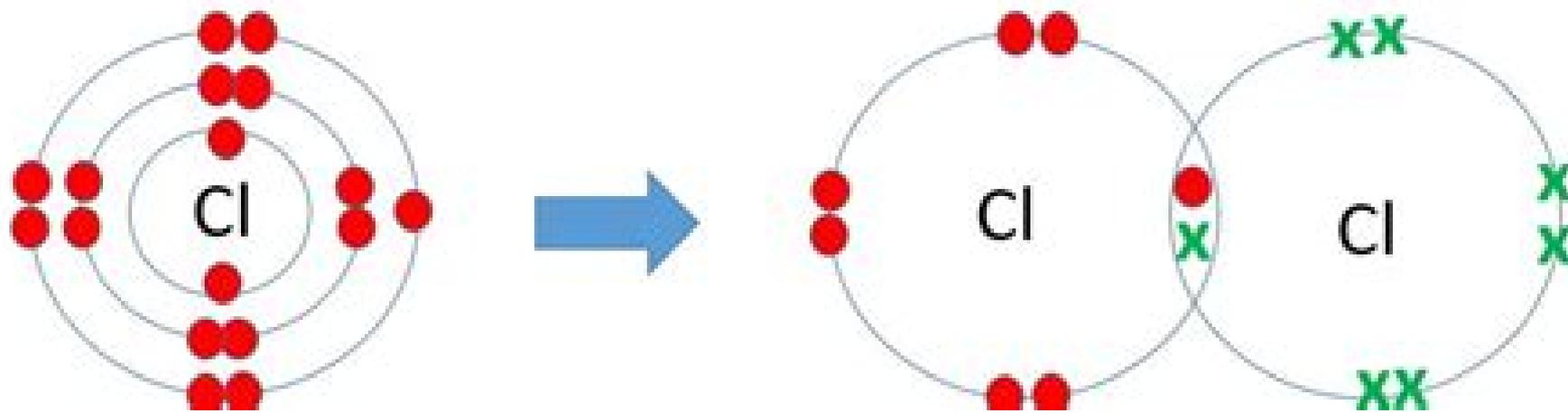


# How to illustrating a covalent bond in Chlorine ( $\text{Cl}_2$ )

1. Draw the electronic structure of an atom of Chlorine
2. How many electrons does it have in its outer shell
3. How many electrons does it need to gain to become stable, (they need to share this number of electrons to gain that many electron, this is also the number of covalent bonds it can form).
4. Draw the atoms so that their outer shell of electron is over lapping draw the correct number of shared electrons in the overlapping region.
5. Draw in the rest of the electrons outside the overlapping regions, both atoms should not have a total of 8 electrons.
6. Note that the original number of electrons in the outer shell of each atom has not changed.



# Covalent Bonding Core Teaching Resources Test Answer

**M Planty**



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