



## Student Exploration: Balancing Chemical Equations

(This document is intended for printing.)

**Essential Question:** How do we balance chemical equations? What is the relationship between the number of atoms of each element on the reactant side and the number of atoms of each element on the product side?

**What You Will Do:** You will use a chemical equation to represent a chemical reaction.

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### What You Will Do

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Example:  $2H_2 + O_2 \rightarrow 2H_2O$

1. Look at the chemical equation  $2H_2 + O_2 \rightarrow 2H_2O$ . The coefficients tell you how many molecules of each substance are involved in the reaction. The subscripts tell you how many atoms of each element are in each molecule.

Reactants:  $2H_2$        $O_2$       Products:  $2H_2O$

2. If a chemical equation is balanced, the number of atoms of each element must be the same on both sides of the equation. (The number of atoms of each element must be the same on both sides of the equation.)

Reactants:  $2H_2$        $O_2$       Products:  $2H_2O$   
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Example: How many molecules of each substance are involved in the reaction? How many atoms of each element are in each molecule?

### What You Will Do

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### What You Will Do



# Explore Learning Stoichiometry Answers

**M Mark**



**Explore Learning Stoichiometry Answers:**

## Unveiling the Magic of Words: A Overview of "**Explore Learning Stoichiometry Answers**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Explore Learning Stoichiometry Answers**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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