

NCERT Solution for Class 12
Chemistry
Chapter 3 - Chemical Kinetics

Intext Exercise

1. For the reaction $R \rightarrow P$, the concentration of a reactant changes from 0.03 M to 0.02 M in 25 minutes. Calculate the average rate of reaction using units of time both in minutes and seconds.

Ans: The average rate of the reaction can be calculated by dividing the change in the rate of decreasing the rate of reactant by the time taken. This is given below:

$$\text{Average rate} = - \frac{d[R]}{dt}$$

This can be written as:

$$\text{Average rate} = - \frac{[R]_2 - [R]_1}{t_2 - t_1}$$

$$R_1 = 0.03 \text{ M}$$

$$R_2 = 0.02 \text{ M}$$

$$t_2 - t_1 = 25 \text{ min}$$

Putting the values, we get:

$$\text{Average rate} = - \frac{0.02 - 0.03}{25}$$

$$\text{Average rate} = - \frac{-0.01}{25} = 4 \times 10^{-4} \text{ M min}^{-1}$$

So, the average rate in minutes will be $4 \times 10^{-4} \text{ M min}^{-1}$

Now, to find the average rate in seconds we have to divide the above answer by 60.

So, the answer will be:

$$\text{Average rate} = - \frac{4 \times 10^{-4}}{60} = 6.66 \times 10^{-6} \text{ M s}^{-1}$$

Therefore, the average rate in seconds will be $6.66 \times 10^{-6} \text{ M s}^{-1}$

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Table of Contents Chapter 4 Assessment Chemistry Answers Pearson

1. Understanding the eBook Chapter 4 Assessment Chemistry Answers Pearson
 - The Rise of Digital Reading Chapter 4 Assessment Chemistry Answers Pearson
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Assessment Chemistry Answers Pearson
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 4 Assessment Chemistry Answers Pearson
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Assessment Chemistry Answers Pearson
 - Personalized Recommendations
 - Chapter 4 Assessment Chemistry Answers Pearson User Reviews and Ratings
 - Chapter 4 Assessment Chemistry Answers Pearson and Bestseller Lists

5. Accessing Chapter 4 Assessment Chemistry Answers Pearson Free and Paid eBooks
 - Chapter 4 Assessment Chemistry Answers Pearson Public Domain eBooks
 - Chapter 4 Assessment Chemistry Answers Pearson eBook Subscription Services
 - Chapter 4 Assessment Chemistry Answers Pearson Budget-Friendly Options
6. Navigating Chapter 4 Assessment Chemistry Answers Pearson eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 4 Assessment Chemistry Answers Pearson Compatibility with Devices
 - Chapter 4 Assessment Chemistry Answers Pearson Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 4 Assessment Chemistry Answers Pearson
 - Highlighting and Note-Taking Chapter 4 Assessment Chemistry Answers Pearson
 - Interactive Elements Chapter 4 Assessment Chemistry Answers Pearson
8. Staying Engaged with Chapter 4 Assessment Chemistry Answers Pearson
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 4 Assessment Chemistry Answers Pearson
9. Balancing eBooks and Physical Books Chapter 4 Assessment Chemistry Answers Pearson
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 4 Assessment Chemistry Answers Pearson
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 4 Assessment Chemistry Answers Pearson
 - Setting Reading Goals Chapter 4 Assessment Chemistry Answers Pearson
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 4 Assessment Chemistry Answers Pearson
 - Fact-Checking eBook Content of Chapter 4 Assessment Chemistry Answers Pearson
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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