

## Covalent Bonding Worksheet

Covalent bonding occurs when two or more **NON\_METALS** share electrons, attempting to attain a stable octet (8 outer electrons) in their outer shell for at least part of the time. Draw a Lewis dot diagram for each element listed. Circle the unpaired electrons that will be shared between the elements.

1.) $H_2$ hydrogen is diatomic: $H + H \longrightarrow \overset{\cdot\cdot}{\underset{\cdot\cdot}{H}} \overset{\cdot\cdot}{\underset{\cdot\cdot}{H}} \longrightarrow H-H \quad \text{Single Bond}$
2.) $F_2$ fluorine is diatomic: $F + F \longrightarrow \overset{\cdot\cdot}{\underset{\cdot\cdot}{\underset{\cdot\cdot}{F}}} \overset{\cdot\cdot}{\underset{\cdot\cdot}{\underset{\cdot\cdot}{F}}} \longrightarrow$
3.) $O_2$ oxygen is diatomic: $O + O \longrightarrow \quad \quad \quad \longrightarrow O=O \quad \text{Double Bond}$
4.) $N_2$ nitrogen is diatomic. Is this a triple bond? $N + N$
5.) $BF_3$ you need 3 fluorine atoms here: $B + F$
6.) Ammonia $NH_3$ hint: how many hydrogen atoms are needed? $N + H$
7.) Carbon dioxide $CO_2$ $O + C + O$
8.) Methane $CH_4$ careful here 4 hydrogen atoms needed $C + H$
9.) Dihydrogen monoxide: the most dangerous substance on the planet. It has killed more people than any other substance known to mankind!! $H + O + H$
10.) $SO_2$ hint: one pair of electrons from sulfur must be slit up for this one to work. $O + S + O$

# Chapter 8 Covalent Bonding Worksheet Answers

## Fruitypiore

**WJ Hussar**



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