

## Section 12.1

### OBJECTIVE 1

Write complete balanced redox reactions that are indicated.

1. a.  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$  (acidic solution)
- b.  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$  (basic solution)
- c.  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$  (neutral solution)
2. a.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (acidic solution)
- b.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (basic solution)
- c.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (neutral solution)
3. a.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (acidic solution)
- b.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (basic solution)
- c.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (neutral solution)
4. a.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (acidic solution)
- b.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (basic solution)
- c.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (neutral solution)
5. a.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (acidic solution)
- b.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (basic solution)
- c.  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$  (neutral solution)

6. Write the net ionic equation for the following balanced equation:



a. Balance the equation:



b. Write the net ionic equation for  $\text{Mn}^{2+}$  and  $\text{H}_2\text{O}_2$ .

c. Is  $\text{Mn}^{2+}$  an oxidizing agent or a reducing agent? Is  $\text{H}_2\text{O}_2$  an oxidizing agent or a reducing agent?

d. Write a balanced equation for  $\text{Mn}^{2+}$  and  $\text{H}_2\text{O}_2$  in acidic solution.

e. Write a balanced equation for  $\text{Mn}^{2+}$  and  $\text{H}_2\text{O}_2$  in basic solution.

# Chapter 9 Review Stoichiometry Answers Section 1

**Jicheng Xie**



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