

Section 12.1

OBJECTIVE 1

Write complete balanced redox reactions that are indicated.

1. $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$ (acidic solution)
 - (a) identify the oxidizing and reducing agents
 - (b) balance the reaction in acidic solution
 - (c) identify the oxidized and reduced species
2. $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$ (basic solution)
 - (a) identify the oxidizing and reducing agents
 - (b) balance the reaction in basic solution
 - (c) identify the oxidized and reduced species
3. $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$ (neutral solution)
 - (a) identify the oxidizing and reducing agents
 - (b) balance the reaction in neutral solution
 - (c) identify the oxidized and reduced species
4. $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$ (acidic solution)
 - (a) identify the oxidizing and reducing agents
 - (b) balance the reaction in acidic solution
 - (c) identify the oxidized and reduced species

5. Write the net ionic equation for the following redox reaction:



a. Balance the equation:



b. Write the net ionic equation for the reaction.

c. Is this a redox reaction? If so, identify the oxidizing and reducing agents.

d. Write the balanced equation for the reaction in acidic solution.

e. Write the balanced equation for the reaction in basic solution.

Chapter 9 Stoichiometry Section 1 Answers

LM Reder



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