

# Chapter 10 mole concept advance additional notes

## What is mole?








This is an atom of carbon-12.  
It is chosen as the standard atom.  
Its  $A_r$  is taken as 12. Then other atoms are compared with it.



This is a magnesium atom.  
It is twice as heavy as a carbon-12 atom, so its  $A_r$  is 24.  
So it follows that ...

A mole of a substance is defined as the number of atoms in an element

### More examples of moles

<p>Sodium is made of single sodium atoms. Its symbol is Na. Its <math>A_r</math> is 23.</p> 	<p>Iodine is made of iodine molecules. Its formula is <math>I_2</math>. Its <math>M_r</math> is 254.</p> 	<p>Water is made of water molecules. Its formula is <math>H_2O</math>. Its <math>M_r</math> is 18.</p> 
 <p>This is 23 grams of sodium. It contains <math>6.02 \times 10^{23}</math> sodium atoms, or 1 mole of sodium atoms.</p>	 <p>This is 254 grams of iodine. It contains <math>6.02 \times 10^{23}</math> iodine molecules, or 1 mole of iodine molecules.</p>	 <p>The beaker contains 18 grams of water, or <math>6.02 \times 10^{23}</math> water molecules, or 1 mole of water molecules.</p>

### Finding the mass of a mole

In order to find the mass of any substance, you must follow these steps:

- Write down the symbol or formula of the substance
- Find its  $M_r$  or  $A_r$ .
- Express that mass in grams(g)

Examples:

Substance	Symbol or formula	$A_r$	$M_r$	Mass of 1 mole
helium	He	He = 4	mass as single atoms	4 grams
oxygen	$O_2$	O = 16	$2 \times 16 = 32$	32 grams
ethanol	$C_2H_5OH$	C = 12 H = 1 O = 16	$2 \times 12 = 24$ $6 \times 1 = 6$ $1 \times 16 = 16$ <hr/> <b>46</b>	46 grams

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**Karin Nielsen-Saines**



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## **Table of Contents Chem11 The Mole Concept Notes 10 Mrowen**

1. Understanding the eBook Chem11 The Mole Concept Notes 10 Mrowen
  - The Rise of Digital Reading Chem11 The Mole Concept Notes 10 Mrowen
  - Advantages of eBooks Over Traditional Books
2. Identifying Chem11 The Mole Concept Notes 10 Mrowen
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chem11 The Mole Concept Notes 10 Mrowen
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chem11 The Mole Concept Notes 10 Mrowen
  - Personalized Recommendations
  - Chem11 The Mole Concept Notes 10 Mrowen User Reviews and Ratings
  - Chem11 The Mole Concept Notes 10 Mrowen and Bestseller Lists

5. Accessing Chem11 The Mole Concept Notes 10 Mrowen Free and Paid eBooks
  - Chem11 The Mole Concept Notes 10 Mrowen Public Domain eBooks
  - Chem11 The Mole Concept Notes 10 Mrowen eBook Subscription Services
  - Chem11 The Mole Concept Notes 10 Mrowen Budget-Friendly Options
6. Navigating Chem11 The Mole Concept Notes 10 Mrowen eBook Formats
  - ePub, PDF, MOBI, and More
  - Chem11 The Mole Concept Notes 10 Mrowen Compatibility with Devices
  - Chem11 The Mole Concept Notes 10 Mrowen Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chem11 The Mole Concept Notes 10 Mrowen
  - Highlighting and Note-Taking Chem11 The Mole Concept Notes 10 Mrowen
  - Interactive Elements Chem11 The Mole Concept Notes 10 Mrowen
8. Staying Engaged with Chem11 The Mole Concept Notes 10 Mrowen
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chem11 The Mole Concept Notes 10 Mrowen
9. Balancing eBooks and Physical Books Chem11 The Mole Concept Notes 10 Mrowen
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chem11 The Mole Concept Notes 10 Mrowen
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chem11 The Mole Concept Notes 10 Mrowen
  - Setting Reading Goals Chem11 The Mole Concept Notes 10 Mrowen
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem11 The Mole Concept Notes 10 Mrowen
  - Fact-Checking eBook Content of Chem11 The Mole Concept Notes 10 Mrowen
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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