

Name Morgan WhiteSection 1B1

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1. 1.75 moles of  $\text{H}_2\text{O}_2$  were placed in a 2.50 L reaction chamber at  $307^\circ\text{C}$ . After equilibrium was reached, 1.20 moles of  $\text{H}_2\text{O}_2$  remained. Calculate the equilibrium constant,  $K_c$ , for the reaction.



2. Phosgene,  $\text{COCl}_2$ , a poisonous gas, decomposes according to the equation



Calculate  $K_p$  for this reaction if  $K_c = 0.083$  at  $900^\circ\text{C}$ .

$$K_p = K_c(\text{PT})^{\Delta n}$$

3. Calculate  $K_c$  for the reaction



given that the concentrations of each species at equilibrium are as follows:  $[\text{HI}] = 0.85 \text{ mol/L}$ ,  $[\text{I}_2] = 0.60 \text{ mol/L}$ ,  $[\text{H}_2] = 0.27 \text{ mol/L}$ .

4. Consider this reaction at equilibrium:



If the volume of the system is expanded at constant temperature, what change will occur in the position of the equilibrium?

# Chemistry 12 Equilibrium Lab Report Answers

**JR Anderson**



## **Chemistry 12 Equilibrium Lab Report Answers:**

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