

## Chapter Test II

### A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

#### Column I

1. An unbalanced chemical equation for a reaction that obeys the law of conservation of matter.
2. The number of protons in the nucleus of an atom.
3. The number of electrons in a neutral atom.
4. The ratio of the actual yield to the theoretical yield, expressed as a percent.
5. An unbalanced chemical equation for a reaction that does not obey the law of conservation of matter.
6. An unbalanced chemical equation that obeys the law of conservation of matter.

#### Column II

- a. percent yield
- b. limiting reagent
- c. theoretical yield
- d. stoichiometry
- e. balanced equation
- f. atomic number
- g. percent reagent

### B. Multiple Choice

Choose the best answer for each question.

1. How many molecules of  $\text{H}_2\text{O}$  are produced when 100 g of  $\text{H}_2$  reacts with 100 g of  $\text{O}_2$ ?
  - a. 1000 molecules
  - b. 10000 molecules
  - c. 100000 molecules
  - d. 1000000 molecules
2. Which of the following is not a balanced chemical equation?
  - a.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
  - b.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2 + \text{O}_2$
  - c.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2 + \text{O}_2$
  - d.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2 + \text{O}_2$
3. The ratio of the actual yield to the theoretical yield, expressed as a percent, is called the
  - a. percent yield
  - b. limiting reagent
  - c. theoretical yield
  - d. stoichiometry

# Chemistry Chapter 12 Stoichiometry Test

**Thomas Evangelist**



## Chemistry Chapter 12 Stoichiometry Test:

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