

Chapter Test II

A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct answer in the blank.

Column I

- _____ 1. The volume of a fixed amount of gas varies inversely as the pressure of the gas, provided the temperature is constant.
- _____ 2. The volume of a fixed amount of gas varies directly as the absolute temperature of the gas, provided the pressure is constant.
- _____ 3. The volume of a fixed amount of gas varies directly as the absolute temperature of the gas, provided the pressure is constant.
- _____ 4. The volume of a fixed amount of gas varies directly as the absolute temperature of the gas, provided the pressure is constant.
- _____ 5. A gas exerts pressure on the walls of its container because of the collisions of the gas molecules with the walls.
- _____ 6. The average kinetic energy of gas molecules is proportional to the absolute temperature of the gas.
- _____ 7. The average kinetic energy of gas molecules is proportional to the absolute temperature of the gas.
- _____ 8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
- _____ 9. The average kinetic energy of gas molecules is proportional to the absolute temperature of the gas.
- _____ 10. The average kinetic energy of gas molecules is proportional to the absolute temperature of the gas.

Column II

- a. Boyle's law
- b. Charles' law
- c. Avogadro's law
- d. Gay-Lussac's law
- e. Boyle's law
- f. Boyle's law
- g. Charles' law
- h. Boyle's law
- i. Charles' law
- j. Boyle's law
- k. Boyle's law
- l. Boyle's law
- m. Boyle's law
- n. Boyle's law
- o. Boyle's law
- p. Boyle's law
- q. Boyle's law
- r. Boyle's law
- s. Boyle's law
- t. Boyle's law
- u. Boyle's law
- v. Boyle's law
- w. Boyle's law
- x. Boyle's law
- y. Boyle's law
- z. Boyle's law

B. Multiple Choice

Choose the correct answer and write its letter in the blank.

- _____ 1. According to Boyle's law, if the volume of a fixed amount of gas is doubled, the pressure of the gas will be _____ of its original value.
 - a. one-fourth
 - b. one-half
 - c. two-thirds
 - d. twice

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Albert Kaskel, 1994-04-19 *Chemistry and Our Universe*, In the first of two lectures on the properties of gases review the basic equations that describe their behavior Learn the history of Boyle's law Gay Lussac's law Charles's law and Avogadro's law Then use these four expressions to derive the celebrated ideal gas law [Teacher's Guide for Gases and "airs"](#)
Elementary Science Study (Education Development Center), 1987

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